

## Zumdahl Chemistry 8th Edition Solutions

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Properties of water - Wikipedia

An acid is a molecule or ion capable of donating a proton (hydrogen ion H<sup>+</sup>) (a Brønsted-Lowry acid), or, alternatively, capable of forming a covalent bond with an electron pair (a Lewis acid). The first category of acids are the proton donors, or Brønsted-Lowry acids. In the special case of aqueous solutions, proton donors form the hydronium ion H<sub>3</sub>O<sup>+</sup> and are known as Arrhenius acids.

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Water has a very high specific heat capacity of 4.1814 J/(g·K) at 25 °C – the second highest among all the heteroatomic species (after ammonia), as well as a high heat of vaporization (40.65 kJ/mol or 2257 kJ/kg at the normal boiling point), both of which are a result of the extensive hydrogen bonding between its molecules. These two unusual properties allow water to moderate Earth's ...

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