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class 10 $\text{NH}_4\text{OH}[\text{aq}] \rightarrow \text{NH}_4^+ + \text{OH}^-$ monobasic acid $\text{HCl} + \text{H}_2\text{O} \rightarrow \text{H}_3\text{O}^+ + \text{Cl}^-$ $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$ dibasic acid $\text{H}_2\text{SO}_4 + 2\text{H}_2\text{O} \rightarrow 2\text{H}_3\text{O}^+ + \text{SO}_4^{2-}$ $\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow \text{NaHSO}_4 + \text{H}_2\text{O}$ $2\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow \text{Na}_2\text{SO}_4 + 2\text{H}_2\text{O}$ tribasic acid $\text{H}_3\text{PO}_4 + 3\text{H}_2\text{O} \rightarrow 3\text{H}_3\text{O}^+ + \text{PO}_4^{3-}$ $\text{NaOH} + \text{H}_3\text{PO}_4 \rightarrow \text{NaH}_2\text{PO}_4 + \text{H}_2\text{O} \dots$

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There was a typo..Now the correct question is; The total number of factors of the number $N = 4^{11} \times 14^5 \times 11^2$? The Options are (A) 405 (B) 504 (C) 620 (D) 528 This was asked by FTTI-JEE for 8th standard...

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experts at BYJU'S after thorough understanding and research on the concepts included in the chapter Atomic Structure and Chemical Bonding. Studying this solution will boost your confidence immensely and help you face the exam confidently.

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ICSE VIII Mathematics Simple and Compound Interest a certain sum of money invested at the rate of 5%per annum compoud interest,the interest compounded annually.if the difference between the

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interests of third year and first is RS102.50.find the sum.

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Selina Concise Chemistry Solution for ICSE Class 10 Chapter 3 Acids, Bases and Salts Selina Publishers for class 10 Chemistry chapter 3 talks about the definition and properties of acids and bases. The chapter further deals with different reactions and students mainly learn how acids and bases react in different circumstances.

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...

This compound gives a precipitate with silver nitrate corresponding to only one Cl - ion and the total number of ions, in this case, is 2. Hence, we can formulate it as $[\text{CoCl}_2(\text{NH}_3)_4]\text{Cl}$. Werner's Theory and Isomerism. Werner turned his attention towards the geometrical arrangements of the coordinated groups around the central cation.

Werner's Theory:

Introduction, Postulates, Limitations ...

Let's understand the rate of a chemical reaction better. Consider a reaction in which the total volume of the

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system remains constant. Let R be the reactants and P be the products i.e. $R \rightarrow P$. Thus, one mole of reactant R produces one mole of product P . Let the initial concentration of R be $[R]_0$ and the concentration of P be $[P]_0$ at time $t = 0$.

Rate of a Chemical Reaction:
Definition, Equations,
Videos ...

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