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The Ultimate Chemical Equations Greek, or German names. Ag Sn w (ferrum) (natrium) (argentum) (stannum) (plumbum) (wolfram) iron sodium silver tin lead tungsten (cuprum) (kalium) (hydrargyrum) (stibium) (aurium) copper potassium mercury antimony gold Symbols are used as a sort of shorthand in writing the names of elements.

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No reaction 8. Calcium metal is added to a solution of nitrous acid. $\text{Ca}(s) + 2\text{HNO}_2(aq) \rightarrow \text{Ca}(\text{NO}_2)_2(aq) + \text{H}_2(g)$ 9. A pea-sized piece of lithium is added to water. $2\text{Li}(s) + 2\text{HOH}(l)$

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$2\text{LiOH}(\text{aq}) + \text{H}_2(\text{g})$ 10. A solution of iron(III) chloride is poured over a piece of platinum wire. No reaction

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Ultimate equation answers

Re-typed from The Ultimate Chemical Equations Handbook by Hague and Smith Ternary Nomenclature: Acids and salts Containing Halogens and/or Oxygen 1. The halogens, with their

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variable oxidation numbers, allow for a great variety of compounds.

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Scientific, Inc. Correlating questions (ii) from: Rodney Schmitz (2007) Moore County Schools, NC For each of the following eight reactions, in part (i) write a BALANCED equation and in part (ii) answer the question about the reaction.

**Key to Reaction Prediction
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g/Write balanced . Title:

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tion. The equations are of mixed types. This Equations section is worth 15 points and is equal to 15 per. cent of the free-response grade. It is for this section of the exam that this handbook was written. All AP equations "work." This means that in every case, a reaction will occur—there are no "no reac-

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Cations. Come from metals that lose electrons in order to become isoelectronic with a noble gas. Anions. Come from nonmetals that gain electrons to become isoelectronic with a noble gas. Oxidation. The process of losing electrons to become isoelectronic with a noble gas.

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PHYSICAL AND CHEMICAL
PROPERTIES Avinash Gupta,
Ph.D. Senior Principal
Chemical Engineer Chevron
Lummus Global Bloomfield, NJ

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1.1 MOLAR GAS CONSTANT 1.2
1.2 ESTIMATION OF CRITICAL
TEMPERATURE FROM EMPIRICAL
CORRELATION 1.2 1.3 CRITICAL
PROPERTIES FROM GROUP-
CONTRIBUTION METHOD 1.3 1.4
REDLICH-KWONG EQUATION OF
STATE 1.5 1.5 P-V-T
PROPERTIES OF ...

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Dekker, probably the
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