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Physics FAST FILE Chapters 16–20 ...

1. 2. Li: K: O: S: 3. 4. Na: Cl

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Teaching Transparency Worksheet Atomic  
And

Date Name TEACHING TRANSPARENCY  
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the shape of an s orbital? Class 16 Use  
with Chapter 5, Section 5.2 2. What is the  
relationship between the size of an s

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orbital and the principal energy level in

MATH SKILLS TRANSPARENCY  
WORKSHEET 6 Using the Periodic ...

TEACHING TRANSPARENCY

WORKSHEET Ionic Bonds Use with  
Chapter 7, Section 7.2 1. How many  
valence electrons does a neutral

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magnesium (Mg) atom have? 2. What is the charge on a magnesium ion? What does magnesium have to do to form such an ion, and why does it tend to do so? 3. How many valence electrons does a single neutral chlorine (Cl) atom have? 4.

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6 41 1. Identify the number of valence  
electrons in each of the following  
elements. a. Ne e. O b. K f. Cl c. B g. P d.  
... Atomic radii cannot be measured  
directly because the electron cloud  
surrounding the nucleus does not have a

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clearly defined a. charge. b. mass. c.

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16 Atomic Orbitals Use with ...

Answers to: "Atomic and Ionic Radii" –

Transparency 21 1. The size of atoms increases down a group in the periodic table. 2. Li:  $1s^2 2s^1$  K:

$1s^2 2s^2 2p^6 3s^2 3p^6 4s^1$  O:  $1s^2 2s^2 2p^4$  S:

$1s^2 2s^2 2p^6 3s^2 3p^4$  K has a greater number of energy levels than Li, which explains

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the difference in size between the 2 atoms  
The same is true for O and S. 3.

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Chemistry: Matter and Change • Chapter

6 35 1. What are the four sections, or

blocks, of the periodic table? 2. What does

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each block represent? 3. What do elements in the s-block have in common? 4. What is the valence electron configuration of each element in group 1? 5.

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Write an equation showing the relationship between an atom's atomic number and its mass number. 8. Lithium has two isotopes: lithium-6 and lithium-7. Draw a diagram, like those shown on the transparency, for each lithium isotope. Label the protons, electrons, neutrons, and

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electron cloud in each diagram. Teaching  
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Groups and Periods of ...

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Science: Geology, the Environment, and  
the Universe Teaching Transparency

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Name Class Date 1. How can you determine the number of protons in the nucleus of an atom of any of the elements listed in the table? 2. Which element has 14 protons in the nuclei of its atoms? 3.

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Chemistry: Matter and Change • Chapter

5 9 1. What is the shape of an s orbital? 2

What is the relationship between the size  
of an s orbital and the principal energy

level in which it is found? 3. What is the

shape of a p orbital? How many p orbitals

are there in a sublevel? 4. How many

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electrons can each ...

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## UNIT 2 RESOURCES Composition of Earth

atomic mass properties periodic law  
atomic number Henry Moseley Dmitri  
Mendeleev The table below was developed  
by John Newlands and is based on a

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relationship called the law of (1) repeated every (2) element (3) According to this law, the properties of the elements elements. Thus, for example, element two and have similar properties.

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