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Mastery In the space at the left,
write true if the statement is true; if
the statement is false ...**

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138 Study Guide for An Introduction to Chemistry stoichiometry. This section shows how to do equation stoichiometry problems for which you are asked to convert from mass of one substance in a given chemical reaction to the corresponding mass of another substance participating in the same reaction. For a related section, see Equation Stoichiometry Problems with Mixtures on our Web site.

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**Chemistry Chapter 11 4
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Chapter 11 Presentation. chapter 11: stoichiometry. Stoichiometry is how a chemist determines the amount of each reactant present at the start of a chemical reaction and how much of a product can form. The solution to every stoichiometric problem requires a balanced chemical equation.

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Name CHAPTER STUDY GUIDE

Date Class Stoichiometry Section

11.1 What is stoichiometry? In your textbook, read about stoichiometry

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and the balanced equation.**

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Chemistry RHS – Mr. Moss 1. Define the following: a. Stoichiometry-the study of the quantitative relationships between the amounts of reactants used and the products formed by a chemical reaction. b. Mole –The SI unit used to measure the amount of a substance that contains 6.02×10^{23} atoms of that substance.

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Chemistry RHS Mr. Moss
[LINK] Chapter 12 Stoichiometry
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GUIDE AND ANSWERS Study Guide
- Chapter 11 – Stoichiometry
Section 11.1 What is stoichiometry?
1. true 2. true 3. false 4. true 5. true
6. 2, 2, 64.10 7. 3, 3, 96.00 8. 2, 2,
88.02 9. 4, 4, 72.08 10. methanol and
oxygen gas 11. carbon dioxide and
water 12. 160.10 g 13. 160.10 g 14.

Chapter 11 Study Guide Chemistry
Stoichiometry Answer Key
Study Guide Page 1 Chemistry
2102A To fulfill the objectives of
this unit, students should complete
the following: Reading for this unit:
Chemistry Chapter 2: Introduction:
page 42 Section 2.1: pages 43 - 46

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**Section 2.2: pages 47 - 54 Section
2.3: pages 55 - 65 Section 2.4:
pages 66 - 74 Chapter 3: Section
3.2: pages 87 - 89 Section 3.3:
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Stoichiometry (Rough outline of the
chapter, please use the book, notes
& homework to study.) 11.1
Defining Stoichiometry Vocab •
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Using Balanced Equations •
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Molecules • Number of Moles •
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Stoichiometry Section 11.1 What is stoichiometry? 1. true 2. true 3.

false 4. true 5. true 6. 2, 2, 64.10 7. 3, 3, 96.00 8. 2, 2, 88.02 9. 4, 4, 72.08

10. methanol and oxygen gas 11.

carbon dioxide and water 12. 160.10 g 13. 160.10 g 14. They are equal.

15. A mole ratio is a ratio between ...

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