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4.2 For some hypothetical metal, the

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equilibrium number of vacancies at 900 C is $2.3 \times 10^{25} \text{ m}^{-3}$. If the density and atomic weight of this metal are 7.40 g/cm^3 and 85.5 g/mol , respectively, calculate the fraction of vacancies for this metal at 900 C.

Solution This problem is solved using two steps: (1) calculate the total number of lattice sites in silver, N , using

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2.2 Silicon has three naturally occurring isotopes: 92.23% of ^{28}Si , with an atomic weight of 27.9769 amu, 4.68% of ^{29}Si , with an atomic weight of 28.9765 amu, and 3.09% of ^{30}Si , with an atomic ...

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2–8. 2 In order to plate a steel part having a surface area of 1250 cm with a. 0.005 cm-thick layer of nickel: (a) How many atoms of nickel are required?

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