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The solubility of reduced Mn(II) is orders of magnitude greater than the solubility of Mn(IV). Therefore, oxidation of Mn(II) to Mn(IV) results in precipitation of manganese oxide solids. Since these solids can exist in several oxidation states (e.g., Mn(III) and Mn(IV)) they are often referred to as $MnO_x (s)$ (Brandhuber et al., 2013).

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The relative solubility of these solutions is considered a factor when determining the appropriate injectate. Highly soluble steroids such as betamethasone sodium phosphate-acetate are rapidly absorbed and pose a lower risk for connective tissue injury, such as tendon rupture, fat atrophy, and muscle wasting.

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(10) Science concepts. The student understands and can apply the factors that influence the behavior of solutions. The student is expected to: (A) describe the

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unique role of water in solutions in terms of polarity; (B) apply the general rules regarding solubility through investigations with aqueous solutions;

**Pharmaceuticals | Solvay
2010 SCORING GUIDELINES
Question 1 (10 points) Several reactions are carried out using AgBr, a cream-colored silver salt for which the value of the solubility-product constant, K_{sp} , is 5.0×10^{-13} at 298 K. (a) Write the expression for the solubility-product constant, K_{sp} , of AgBr. $K_{sp} = [Ag^+][Br^-]$**

**19 TAC Chapter 112,
Subchapter C**

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CYPHOS® 442W is an aqueous solution of a phosphonium salt with high thermal stability and solubility in many common solvents. It has demonstrated use in additives in ethylene-vinyl acetate copolymers in rubber, curable thermosetting, extraction applications from waste and recyclability, denitrogenation of diesel fuel, and organic ...

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