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Answers  
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Perspective  
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Section 2 Quantum  
Theory and the  
Atom (continued)

Chemistry: Matter  
and Change Science

Notebook 63 Bohr's  
Model of the Atom

Use with pages  
146–148. The

Quantum

Mechanical Model of  
the Atom Use with  
page 149–150.

Classify the

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Electron Configurations  
Answers

characteristics of each series in hydrogen's line spectrum. Include the following information. 1.

2.3: Electron Configurations and the Periodic Table ...  
How many protons, neutrons, and electrons are in atoms of these



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Answers

isotopes? Write the complete electron configuration for each isotope.

Answer. Co has 27 protons, 27 electrons, and 33 neutrons:  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^7$ . I has 53

protons, 53 electrons, and 78 neutrons:  $1s^2 2s^2 2p^6 3s^2 3p^6 3d^{10}$

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4s 2 4p 6 4d 10 5s 2

5p 5.

Answers

Section 5.3 electron configuration

Flashcards | Quizlet

Start studying 5.3

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Configurations and Periodic Properties.

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## Electron study tools. Configurations

### 5.3 Electron Configurations and Periodic Properties

...

The electron configurations of silicon (14 electrons), phosphorus (15 electrons), sulfur (16 electrons), chlorine (17 electrons), and

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Answers

argon (18 electrons)  
are analogous in the  
electron

configurations of  
their outer shells to  
their corresponding  
family members  
carbon, nitrogen,  
oxygen, fluorine,  
and neon,  
respectively, except  
that the principal  
quantum number of  
the outer shell of the

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heavier ...

3.4.2 - Electron  
Configuration  
Notations

problems answer  
the following in the  
space provided 22  
write the electron  
configurations for  
the section 51  
models of the atom  
pages 127132  
chapter 5 electrons

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Configurations

Answers

in ... and atomic  
orbitals 53 electron  
configuration main  
idea a set of three  
rules can be ...  
electrons in atoms  
teacher notes and  
answers chapter 4  
section 1 short  
answer 1 in order ...

5 The Periodic Law

If you submit a  
theoretically correct

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Electron  
Configurations  
Answers

configuration, the Gizmo will give you the actual

configuration.)

Element Atomic

number Electron

configuration Cobalt

(Co) 27 1s<sup>2</sup> 2s<sup>2</sup> 2p<sup>6</sup>

3s<sup>2</sup> 3p<sup>6</sup> 3d<sup>7</sup> 4s<sup>2</sup>

Germanium (Ge) 32

1s<sup>2</sup> 2s<sup>2</sup> 2p<sup>6</sup> 3s<sup>2</sup>

3p<sup>6</sup> 3d<sup>10</sup> 4s<sup>2</sup> 4p<sup>2</sup>

Yttrium (Y) 39 1s<sup>2</sup>

2s<sup>2</sup> 2p<sup>6</sup> 3s<sup>2</sup> 3p<sup>6</sup>

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Electron

3d 10 4s 2 4p 6 4d 2

5s 1 Neodymium

(Nd) 60 1s 2 ...

Think and discuss

Compare the

electron

configurations of ...

For example, when

fluorine gains an

electron, the

electron

configuration of F-

will be 1s 2 2s 2 2p



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### Electron Configurations Answers

6. Summary The arrangement of atoms in the periodic table results in blocks corresponding to filling of the  $ns$ ,  $np$ ,  $nd$ , and  $nf$  orbitals to produce the distinctive chemical properties of the elements in the  $s$  block,  $p$  block,  $d$  block, and  $f$  block,

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Configurations

respectively.

Answers Chapter 4.3 :

Electron

Configuration

Section 5.3

continued Answer

the following

questions. Class

STUDY GUIDE FOR

CONTENT

MASTERY Hund's

rule stable Section

5.3 Electron

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Electron Configurations In your textbook, read about ground-state electron configurations. Use each of the terms below just once to complete the passage. Aufbau principle electron configuration ground-state electron ...

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## Electron Configurations Answers

Weebly

The labeled

transitions (A

through E) represent

an electron moving

between energy

levels. If an electron

at level 1 in a

hydrogen atom

absorbs 10.2 eV of

energy, it moves to

level 2. What

typically happens

next? Study Section

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5.3 of The Cosmic  
Perspective.

## Answers

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orthschools.com

Electron

Configurations Look  
at the orbital filling  
diagram of the  
oxygen atom. •

Each of the three 2p  
orbitals has one  
electron. The  
remaining electron

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Electron

now pairs with an  
electron occupying

one of the 2p

orbitals. Electron

Configurations of

Selected Elements

Element 1s 2s 2p x

2p y 2p z 3s Electron

configuration H 1s1

5 2 Electron

Arrangement In

Atoms Section

Review

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Electron  
Configurations  
Answers

Chapter 5.2 :  
Electron  
Configuration and  
the Periodic Table 1.

Chapter 5.2  
Electron  
Configuration and  
the Periodic  
Table 2.

Objectives  
Describe the  
relationship  
between electrons in  
sublevels and the

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## Electron Configurations

length of each  
period of the  
periodic table.<br  
>Locate and name  
the four blocks of  
the periodic table.

Section 53 – Review  
11 - Free CCNA  
Exam Answers 2020  
Chapter 4.3 :  
Electron  
Configuration 1.  
Section 4-3



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Objectives  
List the total number of electrons needed to fully occupy each main energy level.  
State the Aufbau principle, the Pauli Exclusion principle, and Hund's rule.  
Describe the electron configurations for the atoms of any

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Electron

element using  
orbital notation, elec  
tron-configuration  
notation, and when  
appropriate, noble-  
gas ...

5 2 Electron

Arrangement In

Atoms Section

Review

electron

arrangement in

chapter 4 review

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Electron

arrangement of  
Configurations  
electrons in atoms

section 2 short

answer answer the  
following questions  
in the space

provided 1 d how  
many quantum  
numbers are used to

... states of atoms  
and atomic orbitals

53 electron

configuration main

idea a set of three

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rules can be used

Configurations

Section 5.3 Electron  
Configuration

Section 53 Answers.

Standard ACL:

```
access-list x permit  
host y.y.y.y or
```

```
access-list x permit  
x.x.x.x x.x.x.x
```

Extended ACL:

```
access-list x  
permit/deny
```

```
{service/protocol}
```

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Configurations

Answers

{source network/IP}

{destination

network/IP} {port#}

Named ACL: Ip

access-list extended

NAME Permit x.x.x.x

x.x.x.x Deny x.x.x.x

x.x.x.x Apply ACLs:

ip access-group x

inside/outside on

interface, access-

class class x in/out

on ...

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## Configurations Answers

Start studying Section 5.3 electron configuration. Learn vocabulary, terms, and more with flashcards, games, and other study tools.

Chapter 5.2 :  
*Page 30/39*

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Electron  
Configurations  
Answers  
the Periodic Table  
Section Quiz:

Electron  
Configurations In  
the space provided,  
write the letter of the  
term or phrase that  
best completes each  
statement or best  
answers each  
question. 1. The  
statement that no

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Electron Configurations  
Answers

two electrons in the same atom can have the same four quantum numbers is a restatement of a. Bohr's law. b. Hund's rule. c. principle. d. the Pauli exclusion ...

3.4 Electronic Structure of Atoms (Electron Configurations ...



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Section 5.3 Electron  
Configuration In  
your textbook, read  
about ground-state  
electron

configurations. ...

Circle the letter of  
the choice that best  
completes the

statement or  
answers the

question. 15. The  
electrons in an

atom's outermost

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Electron  
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orbitals are called a.  
electron dots. b.  
quantum electrons.  
c.

3.1: Electron  
Configurations  
(Problems) -  
Chemistry  
LibreTexts

Beryllium's electron  
configuration is  $1s^2 2s^2$ . As an  
exercise, write the

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Answers

electron configuration for the first 20 elements on the periodic table. Then click to check your answers.

(section 3.4.3)

Section 4.8

Arrangement of  
Electrons in  
Sublevels..p91

5.2 Electron  
Arrangement in

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Electron  
Atoms >

CHEMISTRY YOU

Some of the

worksheets for this  
concept are Chapter  
7 electron

configuration and  
the periodic table,  
Periodic table and  
electrons, Atom and  
periodic table word  
search, A periodic  
table of the  
elements at los

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Electron  
Configurations  
Answers

alamos national,  
Atomic structure  
and electron  
configurations  
multiple, Electron  
configuration  
practice answers,  
Section organizing  
the elements 155160  
155, Quantum  
numbers work  
answers.

Electrons in Atoms

*Page 37/39*

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Electron

Configurations

Answer  
1s<sup>2</sup>2s<sup>2</sup>2p<sup>6</sup> for 1 ion

c. Write the configuration of the most stable ion this element makes. 5. Refer only to the periodic table at the top of the review of Section 2 to answer the following questions on periodic trends. In a.

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Answers

Which has the larger radius, Al or In? Ca  
b. Which has the larger radius, Se or Ca? Ca c.?

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