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Thermodynamic

Problem : Given that the free energy of formation of liquid water is  $-237 \text{ kJ / mol}$ , calculate the potential for the formation of hydrogen and oxygen from water. To solve this problem we must first calculate  $\Delta G$  for the reaction, which is  $-2 (-237 \text{ kJ / mol}) = 474 \text{ kJ / mol}$ . Knowing that  $\Delta G = -nFE$  and  $n = 4$ , we calculate the potential is  $-1.23 \text{ V}$ .

## Thermodynamics Problems and Solutions

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The first law of thermodynamics –

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problems and solutions. 1. 3000 J of

heat is added to a system and 2500

J of work is done by the system.

What is the change in internal

energy of the system? Known :

Heat ( $Q$ ) = +3000 Joule. Work ( $W$ ) =

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+2500 Joule . Wanted: the change in  
internal energy of the system

Solution :

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