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$5 \text{ C}_6\text{H}_5\text{COOH}(\text{l}) \rightarrow 5 \text{ CO}_2(\text{g}) + 7.5 \text{ O}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{l})$ and n is equal to 7 minus 7.5, or minus 0.5. Hence for the combustion of one mole of benzoic acid at 25°C $\Delta H = -U \cdot 0.5 \cdot 8.31 \cdot 298 \text{ J} = -U \cdot 2480 \text{ J}$

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Physical Chemistry: A practical approach Enrique Peacock-Lopez Department of Chemistry Williams College Williamstown, MA 01267 September 4, 1995

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A practical - Williams College

Aim The purpose of this experiment is to determine x in the formula $\text{Fe}(\text{NH}_4)_2(\text{SO}_4)_2 \cdot x\text{H}_2\text{O}$ by titration against a standard solution of potassium manganate (VII) (permanganate). Introduction This experiment provides useful experience in preparation for a practical examination. The directions are similar to those that would be given by an examination board.

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