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CHEMICAL EQUILIBRIUM (ICE METHOD)

The introduction to the POGIL is all that is appropriate to my students, as we will be teaching equilibrium without involving calculations of the equilibrium constant. The remainder of the POGIL activity is very math heavy. Throughout this lesson, students will be asking questions (SP1) about the reactions they observe.

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POGIL: Le Chatelier's Principle Why?

MODEL #1: Disrupting ...

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Chiolo; Revised by E. Graham, L. Giloni
and Ellen Kannengieser Edited by Linda
Padwa and David Hanson, Stony Brook
University MODEL 1: Dynamic
Equilibrium Acme Manufacturing has
been restricted to 100 employees in the
building at one time. Throughout the

5_LeChatliers_Principle_POGIL.pdf - Equilibrium and Le ...

POGIL Activities for AP* Chemistry
Sample Activity: Reaction Quotient. How
do you predict which direction a reaction
will proceed to reach equilibrium?

EQUILIBRIUM POGIL ANSWER

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equilibrium. Equilibrium in chemical reactions is dynamic because the forward and reverse reactions are occurring continuously and simultaneously at the same rates. Placing a stress on any equilibrium system, whether it is chemical, biological, societal, environmental, or personal, causes the equilibrium position to change. Le

Equilibrium and Le Chatelier's Principle

Chem 116 POGIL Worksheet - Week 8
Equilibrium Continued Why? When a reaction reaches equilibrium we can calculate the concentrations of all species, both reactants and products, by using information about starting concentrations or pressures and the numerical value of the equilibrium constant. Knowing how to set up and solve equilibrium

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Pogil Chemistry Equilibrium And Le equilibrium system, whether it is chemical, biological, societal, environmental, or personal, causes the equilibrium be disrupted. Le Chatelier's Principle allows us to predict the results that follow from "stressing," or in other words, changing the conditions of a system at chemical equilibrium.

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CHEM 116 POGIL Discussion Fall 2010

UMass Boston 1 CHEMICAL
EQUILIBRIUM (ICE METHOD)

Introduction • Chemical equilibrium occurs when opposing reactions are proceeding at equal rates. • The rate at which the products are formed from the reactants equals the rate at which

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Equilibrium Experiment Activity;
Dynamic Equilibrium POGIL; Chemical
Equilibrium Podcast; LeChatelier's
Principle POGIL; LeChatelier's Principle
Podcast; Supplemental DE Notes B
(Complete above by LeChatlier's Principle
Lab) Stress Management Application;
LeChatelier's Principle WS; LeChatelier's
Principle Moodle Quiz; Equilibrium
Constants Podcast

Chemistry

Equilibrium is reached again with the
precipitation of more NaCl to compensate
for the added Cl⁻ ions by shifting to the
left Over-concentrated KSCN +
Fe₂(NO₃)₃ solution Immersed in boiling
water, the solution slowly turned blue
Fe³⁺(aq) + SCN⁻(aq) FeSCN²⁺(aq)
Removing heat

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Chem 116 POGIL Worksheet - Week 8 Equilibrium Continued

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Equilibrium and Le Chatelier's Principle

MODEL 2: LE CHATELIERS

PRINCIPLE Reactant: Increase (T) causes
the equilibrium to shift to the

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AHS Chemistry Resource Site - Unit 6 - Rates & Equilibrium

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(T) causes the

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edia.org/wiki/Le_Châtelier%27s_principle

In chemistry, Le Châtelier's principle, also
called Chatelier's principle or "The
Equilibrium Law", can be used to predict
the effect of a change in conditions on a ...

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two demos, both about the concept of
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uses the reversible reactions program; the
other uses water with food coloring in two
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