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considered neutral (pH 7). If a solution has more H+ than OH-ions, it is considered acidic and has a pH less than 7. If the solution contains more OH-ions than H+ ions, it is considered basic and has a pH greater than 7. The pH scale is logarithmic. This means that you go up and down the scale, the values change in factors of ten.

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that causes the book and the earth to resist being separated. When two oxygen atoms are linked together in a covalent bond, work must be done to separate them. Anything that has the capacity to do such work must, by definition, have energy (Figure 7.2). 250 Chapter 7 Energy and Chemical Reactions

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Chapter 7 Ionic and Metallic Bonding59 SECTION 7.1 IONS (pages 187–193) This section explains how to use the periodic table to infer the number of valence electrons in an atom and draw its electron dot structure. It also describes the formation of cations from metals and anions from nonmetals. Valence Electrons (pages 187–188) 1.

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Organic Chemistry I Jasperse Some Chapter 7 Quiz-Like Practice, But NOT REQUIRED. Answer key available: 1. How many elements of unsaturation are present for a molecule with formula C

SECTION 7.1 IONS (pages 187–193)

This is the lecture recording for Chemistry & Society, Chapter 7 - Solutions and the pH Scale.

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7.7 Solution Concentration In chemistry, concentration is defined as the abundance of a constituent divided by the total volume of a mixture. All of us have a qualitative idea of what is meant by concentration .

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