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-c Determine the percentage composition of each of the compounds below. 1. KMI'IC)t1 I

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What is the percent composition of KMnO₄ - Answers

(role* Answer: 6. Five ml of ethanol has a mass of 3.9 g, and 5.0 mL of benzene has a mass of 4.4 g. Which liquid is denser? Answer: 7. A sample of iron has the dimensions of 2 cm x 3

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cm x 2 cm. If the mass of this rectangular-shaped object is 94 g, what is the density of iron? Answer:

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compositions (rounded to the nearest
tenth of a percent) of all

Percent Composition Worksheet

The elemental composition of

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potassium permanganate is: - oxygen:
40,52 5. - manganese: 34,80%. -
potassium: 24,68 %. You can use
percent composition to see how much
much flour contributes to the mass of
an overall cake. You can see how
much kool aid powder is in a given
amount of kool aid drink, in

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percentage.

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Mass of Sample $\times 100 =$ Percent
Composition. Total mass in the
sample. So, in the M&M's problem: 26
 $\text{red} \times 100 = 13\%$. 200 total. Example
2: Find the percent composition of

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copper (Cu) in CuBr_2 . Mass of Cu in the compound $\times 100 =$ Percent Composition. Total Mass of CuBr_2 Compound. 63.546 grams $\times 100 = 28.4\%$. 223.354 grams

Chemistry 1 Chapter 7 Flashcards | Quizlet

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Chapter 6: Standard Review

Worksheet 1. $1 \text{ amu} = 1.66 \times 10^{-24} \text{ g}$.

For example, the average atomic mass of sodium is 22.99 amu, which represents the average mass of all the sodium atoms in the world (including all the various isotopes and their relative abundances). So that we will

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be able to use the mass of a

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percentage composition: vapor.

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detailed response to each problem.
Show ALL work. including units,
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applicable equations. Utilize dimensional analysis format whenever possible. 1. Calculate the COMPLETE percentage composition (by mass) of the following:

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**PERCENTAGE COMPOSITION
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COMPOSITION OF HYDRATES

Name A hydrate is an ionic compound with water molecules loosely bonded to its crystal structure. The water is in a specific ratio to each formula unit of the salt. For example, the formula

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$\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ indicates that there are five water molecules for every one formula unit of CuSO_4 . Answer the questions below. 1.

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Solutions. Find the percent compositions of all of the elements in the following compounds: 1) CuBr 2 Copper (II) Bromide Cu: 28.45% Br: 71.55% 2) NaOH Sodium Hydroxide Na: 57.48% O: 39.99% H: 2.53% 3) (NH₄)₂S Ammonium Sulfide N: 41.1% H: 11.8% S: 47.1% 4) N₂S₂ Dinitrogen

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disulfide.

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