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Laboratory 11: Molecular Compounds and Lewis Structures ...

a. type of bonding b. molecular shape c. molecular polarity. for each of the following compounds (construct a table): (1) HBr (3) BaCl₂ (5) Cl₄ (2) SCl₂ (4) NH₃ (6) AlH₃. 2.

Calculate the electronegativity difference and indicate the type of bond for ... LAB: SHAPES OF COVALENT MOLECULES & POLARITY ...

Lab Report: VSEPR Theory and the Shapes of Molecules

Possible Shapes for Molecules with 4 electron groups around the central atom (bond angles

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near 109.5o) Molecular Shape 3D sketch and Lewis structure Tetrahedral A four sided figure. Each side is an equilateral triangle. Trigonal pyramidal A four sided figure. Three sides are isosceles triangles.

Lab 11 Worksheet | Chemistry I Laboratory Manual

Table 1: Molecular Shapes We will be using computer models to verify the VSEPR predictions we make. These models have been validated by experiment, so we will not need to experimentally measure any molecules. In later chemistry courses, you will have the

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opportunity to learn some of the methods that are used to test computational models.

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Lab #2: Molecular Models Work in groups of 3-4, each group uses two model kits. Bring your textbook. Refer to pages 23, 34-43. One of the difficulties of studying molecular bonding is that you cannot see atoms and molecules. It is difficult to visualize the shape of a molecule based on a two-dimensional drawing in a textbook. In this

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Laboratory 11: Molecular Compounds and Lewis Structures Molecular Model Building (3D Models) The 3D structure of molecules is often difficult to visualize from a 2D Lewis structure. In order to understand the true 3D shape of molecules molecular model kits will be used to create 3D models. This will make it easier to see the common

5.03 Lab Molecular Models.PDF - Lab 11
Molecular Models ...

Although you do not need to name the molecular shape for molecules and ions with more than one "central atom", you should be

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able to indicate the molecular geometry about each "central atom." Click here to review VSEPR theory. During lab construct a molecular model, using the kit provided, for each species listed in the tables.

LAB: SHAPES OF COVALENT MOLECULES & POLARITY

119 Lab 11: Molecular Models Introduction Why

can't you play basketball with a football?

The obvious answer is because a football isn't the right shape. A football can't be dribbled and would be very difficult to shoot. On the other hand, a fairly large spherical ball can be easily dribbled and is

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needed to successfully play basketball

Pre-Lab #2: Molecular Models

When is a molecule polar? Change the electronegativity of atoms in a molecule to see how it affects polarity. See how the molecule behaves in an electric field. Change the bond angle to see how shape affects polarity.

Molecule Polarity - Polarity |

Electronegativity | Bonds ...

Constructing Models, Determining Molecular Shapes and Molecular Polarity. Use your

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molecular model kit to construct a three-dimensional model of each of these molecules and polyatomic ions. Sketch a reasonably detailed picture of this model on your Report Form. Rules for Constructing Molecules with the Model Kit

Experiment 11: MOLECULAR GEOMETRY & POLARITY
With the help of a molecular model kit and a computer modeling program, you will be able to visualize a molecule in three-dimensions. In this lab, you will use a computer program within WebAssign that allows molecules to be rotated, just like you could manually rotate

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a model built with a model kit.

Molecule Shapes - Molecules | VSEPR | Lone Pairs - PhET ...

LAB EIGHT Shapes of Molecules 1 Name Lab Partner(s) Section Date Shapes of Molecules

Objective In this experiment you will build models of molecules to help you determine their electron and molecular geometries.

Introduction Electron geometry is the shape of the molecule produced by the position of the electron pairs around the central atoms.

LAB EIGHT - Lake-Sumter State College

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molecular shape and arrangement of atoms. Model Building Basics Molecular model kits vary; therefore, your instructor will explain the particular models that you will use. The kit probably contains balls (used for atoms), sticks (used for single bonds and unshared electron pairs), and springs or curved sticks (used for double and triple bonds).

Molecular Modeling 1 | Chem Lab
Molecular geometry refers to the 3-D shapes of molecules and polyatomic ions.

Lab Activity: Molecular Model Building -

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Bellevue College

Chemical Bonds, Molecular Models, and
Molecular Shapes PRELAB ASSIGNMENT Write an
objective, and answer the following questions
in your laboratory notebook before coming to
lab. Read the entire laboratory write up
before answering the prelab questions. 1.

Chemical Bonds, Molecular Models, and
Molecular Shapes

Post-lab Questions. 1. Without making a
model, describe the electron geometry and
molecular shape of carbon tetrabromide (CBr
4). Would you expect the bonds in this

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molecule to be polar? Would you expect this molecule to be polar overall? Explain. 2. NH_3 and H_2CO each have three bonds about the central atom.

9: Lewis Structures and Molecular Shapes (Experiment ...

A molecular model is far superior to a structural formula when it comes to visualizing atomic arrangement. Compared to molecular formulas and structural formulas, molecular models provide much more information about the true shapes of molecules. In this experiment, you will use

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ball-and-stick models to help you visualize the shapes of molecules.

Molecular Models Lab Instructions

Lab Partner: Lab Section: Lab Report: VSEPR

Theory and the Shapes of Molecules HCN 1.

Lewis Structure 2. Perspective drawing 3.

Number of atoms bonded to central atom 4.

Number of lone electron pairs on central atom

5. Electronic geometry and ideal bond angles

6. Molecular geometry 7. Hybridization of

central atom 8. Molecular polarity CH₃OH 1

Chemical Bonds, Molecular Models and Shapes

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Lab 5 - Molecular Geometry

Explore molecule shapes by building molecules in 3D! How does molecule shape change with different numbers of bonds and electron pairs? Find out by adding single, double or triple bonds and lone pairs to the central atom. Then, compare the model to real molecules!

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The molecular shape is the shape of the atoms

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alone. Large molecules have more than one central atom and therefore more than one shape. The common shapes that we will examine in this lab are linear, bent, trigonal planar, trigonal pyramid, and tetrahedral.

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