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CHAPTER 11 Gases
Gases SECTION 3 SHORT ANSWER
Answer the following
questions in the space
provided. 1. The molar mass
of a gas at STP is the

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density of that gas (a) multiplied by the mass of 1 mol. (c) multiplied by 22.4 L. (b) divided by the mass of 1 mol. (d) divided by 22.4 L. 2. For the expression $V = nRT/P$, which of the following will cause the volume to increase?

Modern Chemistry: Chapter 10
"Physical Characteristics of
...

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Teacher Notes and Answers
Chapter 12 SECTION 1 SHORT
ANSWER 1. c 2. a 3. b 2. a.
alcohol b. water c. the gels
3. The mixture is a colloid.
The properties are
consistent with those

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reported in Table 3 on page 404 of the text. The particle size is small, but not too small, and the mixture

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10 States of Matter - Ms. Agostine's Chemistry Page
Modern Chemistry 76 Quiz
Section Quiz: Gas Volumes and the Ideal Gas Law In the space provided, write the letter of the term or phrase that best completes each sentence or best answers each question. _____ 1. At the same temperature and

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pressure, balloons of equal volume always contain a. equal masses of gas. b. equal numbers of molecules.

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equal volumes of gases at the same temperature and pressure contain equal numbers of molecules $V = kn$
V-volume k-constant n-amount of gas in moles
Standard Molar Volume of a Gas the volume occupied by one mole of a gas at STP =22.414 L

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1. Gases consist of large

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numbers of tiny particles that are far apart relative to their size. 2. Collisions are elastic between particles and the container. 3. Particles are always in motion. No attractive forces between them except near condensation point. 4. No forces of attraction or repulsion between particles. 5.

CHAPTER 10 States Matter

$P_T = P_1 + P_2 + P_3 + \dots$
. where P_T is the total pressure of the mixture, P_1 is the partial pressure of the first gas, P_2 is the partial pressure of the second gas, and so on. The kinetic-molecular theory of

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matter can explain Dalton's law.

mc06se cFMsr i-vi

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CHAPTER 12 REVIEW Solutions
The Periodic Law SECTION 1
SHORT ANSWER Answer the following questions in the space provided. 1. c In the

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modern periodic table,
elements are ordered (a)
according to decreasing
atomic mass. (b) according
to Mendeleev's original
design. (c) according to
increasing atomic number.
(d) based on when they were
discovered.

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...

States of Matter SECTION 3
SHORT ANSWER Answer the
following questions in the
space provided. 1. Match
description on the right to
the correct crystal type on
the left. b ionic crystal
(a) has mobile electrons in
the crystal c covalent

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molecular crystal (b) is
hard, brittle, and
nonconducting

5 The Periodic Law

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CHAPTER 11 REVIEW Gases
SECTION 3 SHORT ANSWER
Answer the following

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questions in the space provided. 1. _____ The molar mass of a gas at STP is the density of that gas (a) multiplied by the mass of 1 mol. (c) multiplied by 22.4 L. (b) divided by the mass of 1 mol. (d) divided by 22.4 L. 2. _____ For the expression , P

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(a) The molar masses and the velocities are inversely proportional. (b) The molar masses and the velocities are directly proportional. (c) The molar masses and the square roots of the velocities are directly proportional. (d) The molar

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masses and the squares of the velocities are inversely proportional.

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Chapter 11 - Gases - An
Introduction to Chemistry
Orbitals of equal energy are
each occupied by one

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electron before any orbital is occupied by a second electron, and all electrons in singly occupied orbitals must have the same spin state.

Modern Chemistry : Section Quizzes with Answer Key ... Chapter 11 Gases Test B Holt Modern Chemistry - best free. Latest for Chapter 11 Gases Test B Holt Modern Chemistry. of Gases SECTION 10-1 SHORT ANSWER Answer. PDF] CHAPTER 11 REVIEW Molecular Composition of Gases SECTION 11-2. modern chemistry chapter 11 section 1 molecular composition of gases answers Examples of Solids Liquids Gas

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Modern Chemistry Section 3 Gases

tion 3) carries them throughout the available space. Such spontaneous mixing of the particles of two substances caused by their random motion is called diffusion. Gases diffuse readily into one another and mix together due to the rapid motion of the molecules and the empty space between the molecules.

CHAPTER 11 REVIEW Gases -
Manasquan Public Schools
Section 11.3 Equation
Stoichiometry and Ideal
Gases Goal: To show how gas-

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related calculations can be applied to equation stoichiometry problems. This section shows how we can combine calculations such as those found in Chapter 10 with the gas calculations described in Section 11.2 to do equation stoichiometry problems that

11 Molecular Composition of Gases - Madison Public Schools

combined gas law SECTION 3
Gas Volumes and the Ideal Gas Law KEY TERMS • Gay-Lussac's law of combining volumes states that the volumes of reacting gases and their products at the same temperature and

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pressure can be expressed as ratios of whole numbers.

- Avogadro's law states that equal volumes of gases at the same

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