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the rates of effusion of gases at the same temperature and pressure are inversely proportional to the square roots of their molar masses Rate of effusion of A / B =  $\sqrt{M_B} / \sqrt{M_A}$

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Chapter 11 - Properties of Solutions . 11.1 Solution Composition . A. Molarity i. liters of. solution moles solute Molarity(M) = B. Mass Percent 1.  $\times 100$  = mass of. solution mass of solute Mass percent. C. Mole Fraction . 1. D. Molality 1. ki ram of solvent moles of solute Molality log = E. Normality 1. liter of solution equivalents Normality = 2.

### Holt's Modern Chemistry Textbook | CourseNotes

The Solutions chapter of this Holt McDougal Modern Chemistry Companion Course helps students learn the essential lessons associated with solutions.

### Mrs. J's Chemistry Page - Lecture Notes

Modern Chemistry 136 Chapter Test Name Class Date Chapter Test A, continued \_\_\_\_14. A pH meter measures the pH of a solution by measuring the a. amount of charge on the hydronium ion. b. color of the solution. c. concentration of the solution. d. voltage difference between two electrodes in the meter's probe. \_\_\_\_15. When you conduct an ...

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