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(ii) A student titrated 25.00 cm³ of a 0.100 mol dm⁻³ solution of hydrofluoric acid, HF (aq), with 0.100 mol dm⁻³ NaOH (aq). Some of his data are presented below.

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(d) A solution of ammonia has a concentration of 0.500 mol dm⁻³. Calculate the pH of the ammonia solution using information from Table 15 of the Data Booklet. State one assumption made. [4]..... (e) A buffer solution is made using 25.0 cm³ of 0.500 mol dm⁻³ nitric acid, HNO₃ (aq), and 325.0 cm³ of 1.00 mol dm⁻³ ammonia solution, NH₃ (aq).

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