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What is the percent composition carbon in NaHCO_3 .

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mass of empty test tube mass of test tube with copper ...

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1. Within bounds of experimental error, use the mole ratio to write the formula of the compound. 2. What is the name of this compound? (Hint: you need a Roman numeral in the name) 3. A compound of calcium and bromine is analyzed in the lab. A 20.0 g sample contains 4.00 g calcium. What is the empirical formula of the compound? 4.

Example Problem: Find the % by mass of oxygen in water.

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Episode Name Example Problem: Find the % by mass of oxygen in water. $(16/18) * 100 = 88.9\%$ Percentage by mass

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Example Problem: Find the % by mass of oxygen in water. Percentage by mass of element in a compound = $(\text{mass of element in 1 mol of compound} \div \text{molar mass of compound}) \times 100 \%$

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