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Physical Pharmacy
Lab Report UKM
Points to Remember
while Performing the
Experiment in a Real**

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Laboratory: Always wear lab coat and gloves when you are in the lab. When you enter the lab, switch on the exhaust fan and make sure that all the chemicals and reagents required for the experiment are available.

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solution: An experiment to illustrate the ...
In the method of B.E.T (Brunauer, Emmett and Teller), adsorption of gas was used to measure the surface area. In this experiment, adsorption of iodine from solution is studied and

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Langmuir equation is used to estimate the surface area of activated charcoal sample.

**Experiment 5
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EXPERIMENT 5
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Introduction The**

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term adsorption is used to describe the fact that there is a greater concentration of the adsorbed molecules at the surface of the solid than in the bulk solution. In general, one uses solid adsorbents of small size and often with surface imperfections such

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**Adsorption Isotherm
(Procedure) :
Physical Chemistry**

...

**CHEM355
EXPERIMENT 9**

Adsorption

**Adsorption is the
accumulation of a
gas or liquid solute
(the adsorbate) on a
surface of a solid or**

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a liquid (the adsorbent) forming a molecular or atomic film. It is different from absorption, in which a substance diffuses into a liquid or solid to form a solution.

**Adsorption Short Report
_DeJanteEaton_C
HEM3521 - Eaton 1**

...

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The maximum adsorption of the modified biochar was observed at pH 2.0 with the sorption capacity of 102.66 mg g⁻¹. The adsorbed zinc–biochar could be effectively regenerated by 0.5 mol L⁻¹ NaOH solution and the adsorption ability

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decreased from
84.16 to 59.75 mg g
?1 in the sixth cycle.

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4 7 Using practical
balance and wessels
for weighing coal,
weigh 6 portions of
activated charcoal,
each portion 5 g.
The accuracy of

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weighing must be with accuracy 0.01 g. 8 Put activated charcoal into numbered flasks with stoppers (1 portion per flask). 9 Plug up the flasks, and shake them. Wait for 20 minutes, the process of adsorption is in

Continuous Fixed-

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Bed Column Study and Adsorption

Modeling ...

**Multi-walled and
single-walled carbon
nanotubes were
used as
nanoadsorbents for
the successful
removal of Reactive
Blue 4 textile dye
from aqueous
solutions. The
adsorbents were**

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characterised by infrared and Raman spectroscopy, N₂ adsorption/desorption isotherms and scanning and transmission electron microscopy. The effects of pH, shaking time and temperature on adsorption capacity were studied.

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**Dynamic and
Equilibrium
Adsorption
Experiments.**

**Journal of Chemical
Education 2005, 82
(6) , 919. DOI:
10.1021/ed082p919.**

**Julia A. Gavin,,
Nanlin Deng,,**

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**Mónica Alcalá, and,
Thomas E. Mallouk.**

Host?Guest

**Chemistry of a
Chiral Cyclohexaned
iamine?Viologen
Cyclophane in
Solution and in the
Solid State.**

**Effect of porous
zinc–biochar
nanocomposites on
Cr(vi ...**

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Regeneration was done by 0.2 M HCl solution and the experiment was repeated for four adsorption-desorption cycles. In first cycle 97.2% removal was achieved and in fourth cycle the removal efficiency of 88.1% was achieved (Figure 12).

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FF lab report:

PRACTICAL 3:

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EXPERIMENT 5
ADSORPTION FROM
SOLUTION Lab
Questions 1. They
report the following

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**amounts of
adsorption Gas**

**Adsorption Isotherm
System is a high
pressure, 10,000 psi,
system designed for
the evaluation of
gas adsorption
isotherms, or the
gas capacity of a
shale or or ions) in
solution.**

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CHEM355
EXPERIMENT 9
Adsorption
In this experiment,
adsorption of iodine

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from solution is studied and Langmuir equation is used to estimate the surface area of activated charcoal sample. Materials iodine solutions (specified in Table 1), 1 % w/v starch solution, 0.1 M sodium thiosulphate solution, distilled water and activated

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charcoal.
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EXPERIMENT 3:
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DeJante Eaton Lab

Partner: Ben

Ingolfsson TA:

Melissa Foxley

February 23, 2016

**INTRODUCTION In
this experiment, the
adsorption of
activated carbon
from an acetic acid
solution was
determined; in**

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addition we were able to use the Langmuir equation to characterize surface area of the activated carbon. This is important in determining the relationship between ...

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Questions 1. They
report the following
amounts of
adsorption Gas
Adsorption Isotherm
System is a high
pressure, 10,000 psi,

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**system designed for
the evaluation of
gas adsorption
isotherms, or the
gas capacity of a
shale or or ions) in
...**

**Exercise 5
DETERMINATION
OF ADSORPTION
ISOTHERM OF
ACETIC ...
EXPERIMENT 5**

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**ADSORPTION OF
ACETIC ACID ON TO
ACTIVATED
CHARCOAL
SUGGESTED**

BACKGROUND

READING Atkins,

P.W., Physical

Chemistry, 6 th ed.,

7 th ed., Oxford

University Press,

Oxford, 1998/9

(Chapter 28) Atkins,

P.W., & Julio de

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**Paula, Physical
Chemistry, 8 th ed.,
Oxford University
Press, Oxford, 2006
(Chapter 25)**

INTRODUCTION

**Activated charcoal
or carbon is widely
used for vapour
adsorption and ...**

**EXPERIMENT 5 -
Adsorption of Acetic
acid on charcoal ...**

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OBJECTIVE 1.) To study the adsorption of iodine from solution 2.) To identify the surface area of activated charcoal sample using Langmuir equation.

Introduction

Adsorption is a process of free moving of gaseous or solutes

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**molecules of a
solution come close
and attach
themselves onto
surface of solid. The
adsorption can be
strong or weak
depends...**

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