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Chapter 21 – Nuclear Chemistry: Part 1 of 9

Honesty On each exam day I am going to give you two examinations, one in chemistry and one in honesty. I hope you will pass them both, but if you must fail one, let it be chemistry, for there are many good people in this world today who cannot pass an examination in chemistry, but there are no good people in the world who cannot pass an examination in honesty.

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Nuclear Chemistry Nuclear Transformations • Rutherford in 1919 performed the first nuclear transformation. • The transmutations are sometimes represented by listing in order, the target nucleus, the bombarding particle, the ejecting particle and the product nucleus. • The above equation becomes: ${}_{14}^{27}\text{Al} + {}_2^4\text{He} \rightarrow {}_{17}^{31}\text{P} + {}_1^1\text{H}$

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CHAPTER 23 NUCLEAR CHEMISTRY 23.1 THE NATURE OF NUCLEAR REACTIONS

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radioactivity - the spontaneous decay of an unstable nucleus with accompanying emission of radiation. nuclide - atom with a specific number of protons and neutrons in its nucleus. ? There are 271 stable nuclides in nature, others are radioactive radionuclide - unstable isotope that undergoes nuclear decay

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General Chemistry II

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Chemistry 1110 – Chapter 5 – Nuclear Chemistry – Practice Problems Page | 5 22. The nuclear reaction shown below is an example of what type of process? ${}^{235}_{92}\text{U} + {}^1_0\text{n} \rightarrow {}^{141}_{54}\text{Xe} + {}^{92}_{38}\text{Sr} + 3{}^1_0\text{n}$ A) fusion B) fission C) translation D) alpha decay E) beta decay 23. Nitrogen-17 is a beta emitter. What is the isotopic notation for Nitrogen-17?

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produced in the radioactive decay? A) C B) B C) N D) F

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640 CHAPTER 23: NUCLEAR CHEMISTRY This is the nuclear binding energy. It's the energy required to break up one tungsten-184 nucleus into 74 protons and 110 neutrons. When comparing the stability of any two nuclei we must account for the fact that they have different numbers of nucleons.

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