

## Essentials Of Chemical Reaction Engineering Solution Manual

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Essentials of Chemical Reaction Engineering is a complete yet concise, modern introduction to chemical reaction engineering for undergraduate students. While the classic Elements of Chemical Reaction Engineering, Fourth Edition, is still available, H. Scott Fogler distilled that larger text into this volume of essential topics for undergraduate students.

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Stoichiometry. We shall set up Stoichiometric Tables using A as our basis of calculation in the following reaction. We will use the stoichiometric tables to express the concentration as a function of conversion. We will combine  $C_i = f(X)$  with the appropriate rate law to obtain  $-r_A = f(X)$ .

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Chemical Reaction Engineering: Fogler & Gurmen

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P3-15 Assuming the reactions to be elementary: 2 Anthracene  $\rightarrow$  Dimer 2 ( ) Dimer A Anthracene  $C_6H_4$   $C_6H_4$   $C_6H_4$  where,  $K_c = k_1 + /k_2$ -Similarly for the second reaction: 3-13 Norbornadiene Quadricyclane ( ) Quadricyclane Norbornadiene Norbornadiene  $C_6H_4$   $C_6H_4$   $C_6H_4$  where,  $K_c = k_1 + /k_2$ -P3-16 The mistakes are as follows: 1. For any reaction, the rate law cannot be written on the basis of the stoichiometric equation.

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