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laboratory reports are developed to interact the findings of research, in such a way that is clear to readers. You need to not forget to include any extra details, which may be helpful for readers.

**Lab 1: Chemical
Equilibrium: Finding
a Constant, K_c**

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***Determination of an
Equilibrium***

***Constant for the Iron
(III) thiocynate***

Reaction Pre-lab

***Assignment Before
coming to lab: •***

***Read the lab
thoroughly. •***

***Answer the pre-lab
questions that
appear at the end of
this lab exercise.***

The questions

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***should be answered
on a separate (new)
page of your lab
notebook. Be sure
to show all***

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equilibrium constant can then be determined for each mixture; the average should be the equilibrium constant value for the formation of the FeSCN^{2+} ion. In Part A of this experiment, you will prepare FeSCN^{2+} solutions of known

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**concentrations,
measure their ...**

***Post Lab Questions
Chemistry***

Experiment 17

Equilibrium

Lab:

***Spectrophotometric
Determination of an
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***Constant.. Post Lab
Question: A student
rushes through the***

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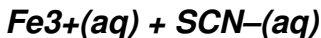
lab. Specifically, he neglects to wipe the cuvettes clean of smudges and fingerprints prior to measuring the absorbances.

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Lab 1: Chemical
Equilibrium: Finding**

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a Constant, K_c The purpose of this lab is to experimentally determine the equilibrium

constant, K_c , for the following chemical reaction:



***iron(III) thiocyanate
thiocyanoiron(III)***

When Fe^{3+} and SCN^{-} are combined,

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equilibrium is

**established between
these two ions and
the**

**Experiment 3
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Constant for ...
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***Lab 5 -
Determination of an
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Constant
CHEM113L:***

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**equilibrium constant
can then be
determined for each
mixture; the average
should be the
equilibrium constant**

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value for the formation of the FeSCN^{2+} ion. In Part A of this experiment, you will prepare FeSCN^{2+} solutions of known concentrations, measure their

Spectrophotometric Determination Of An Equilibrium ...

Post-lab Questions:

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Top. Additional problems to be turned in with the laboratory write-up.

SHOW ALL CALCULATIONS. 1.

In the reaction initially 1.80 mol of A, 4.20 mol of B and 1.00 mol of C are placed in a 3.00 liter container. Analysis shows 4.50 mol of B at equilibrium.

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**Calculate the
equilibrium constant
for the reaction. 2.**

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Science Notes 2: ...

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Answers attended. I

have modified it

some along the

way.) The lab is a

pretty basic look at

causing disruptions

to an equilibrium

and making

predictions - and

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***Determination Of
Equilibrium
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***An equilibrium
constant can then
be determined for
each mixture; the
average should be
the equilibrium
constant value for
the formation of the
FeSCN 2+ ion. In***

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Part A of this experiment, you will prepare FeSCN^{2+} solutions of known concentrations, measure their absorbance at 470 nm, and produce a calibration curve.

***Solved: Lab:
Spectrophotometric
Determination Of An
Equili ...***

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**Previous Answers N
CSUGenChem102La**

bV1 8.POST.09. In

this exercise, you

will calculate the

equilibrium constant

(K eq) and pK eq for

the reaction studied

in Part C. Data will

be needed from

Table C of the Lab 8

in-Lab assignment.

(a) Calculate the [H³

O⁺] for each buffer

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**solution and enter
the value in Table D.
Use the formula ...**

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