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Information: The Gas Constant

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a) How many moles of gas are in his lungs? 0.343 mol b) If the diver swims to the surface without exhaling where the temperature is 32oC and the pressure changes to 100.2 kPa, what will the volume of the air in his lungs be? 8.68 L 3. On planet X, 2.78 moles of a gas takes up 1.85 L under a pressure of 74.1 kPa and a temperature of 201oC.

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Gas (# of moles of gas is constant) Gas # of moles of Gas in container (mol) Volume of Container (L) Temperature of Gas (K) Pressure of Gas (kPa) A 0.3008 14.0 672.0 120.0 B 0.3008 16.0 1280 200.0 C 0.3008 10.0 340.0 85.0 Table 2: Experimental Data for Three Different Samples of Gases That Contain Different Amounts of Gas (# of moles of gas is ...

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