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[2] (b) Marit arranged the values she found in Table 12 into an energy cycle. $C_2H_4(g) + H_2(g) \rightarrow C_2H_6(g)$
 $3 O_2 \rightarrow 2 CO_2(g) + 3 H_2O$
□ 1411 kJ mol⁻¹ □ 1 □ 286 kJ mol⁻¹ □ 1 □ 1560 kJ mol⁻¹ □ 1 (hydrogenation) (1)
Calculate the value for the enthalpy of hydrogenation of ethene from the energy cycle.....

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□ 2 □ Section A Answer all questions.

Write your answers in the boxes provided.

1. Ethanedioic acid is a diprotic acid. A student determined the value of x in the formula of hydrated ethanedioic acid, $\text{HOOC}\cdot\text{COOH}\cdot x\text{H}_2\text{O}$, by titrating a known mass of the acid with a $0.100 \text{ mol dm}^{-3}$ solution of NaOH (aq). 0.795 g of ethanedioic acid was dissolved in distilled water and made up to a total ...

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$4\text{NH}_3(\text{g}) + 5\text{O}_2(\text{g}) \rightleftharpoons 4\text{NO}(\text{g}) + 6\text{H}_2\text{O}(\text{g})$ $\Delta H = -910 \text{ kJ mol}^{-1}$

Describe the effect, if any, of each of the following changes on the equilibrium concentration of nitrogen monoxide in a particular equilibrium mixture, giving a reason in each case. (i) Increasing the pressure, at constant temperature [2]

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Paper 2 □ 2 □

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