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Chapter 9 Stoichiometry Section 1 Answers

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Practice Test Ch3 Stoichiometry (page 3 of 3) 1. d It might be easiest to balance the equation with mostly whole numbers: 2 NH₃ + $\frac{5}{2}$ O₂ → 2 NO₂ + 3 H₂O. The question asks about the amount of oxygen reacting with ONE mole of ammonia, thus cut the $\frac{5}{2}$ (3.5) of oxygen in half to 1.75 2. c Balance: CH

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Given the following equation: $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$ 4 a. What is the value of the coefficient x in this equation? 40.07 g/mol b. What is the molar mass of C_3H_4 ? 2 mol O_2 :1 mol H_2O c. What is the mole ratio ...

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To understand stoichiometry, start with this introduction to the topic. It might also help to review molecules and moles, which includes how chemical formulas work. Ready for another quiz? Here's a quick self-test about the mole. If you'd rather switch gears, see if you know the answers about how chemistry explains the real world.

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equation. Chapter 9 Stoichiometry Table of Contents Chapter 9 Stoichiometry Section 1 Chapter 9 Section 1 Intro to Stoichiometry. STUDY. Flashcards. Learn. Write. Spell. Test. PLAY. Match. Gravity. Created by. Blair12_Armstrong. Key Concepts: Terms in this set (10) Stoichiometry is the branch of chemistry that deals with elements in compounds ...

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