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*Matching Match  
each item with  
the correct  
statement below.  
Match each item  
with the correct  
statement below.  
a. monatomic ion  
f. cation b.*

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acid g. binary  
Answers compound c. base  
h. anion d. law  
of definite  
proportions i.  
polyatomic ion  
e. law of  
multiple  
proportions \_\_\_\_\_  
1.

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*Answers may  
include release  
or absorption of  
energy, change  
in color, change  
in odor,  
formation of a  
gas, or*

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*formation of a  
solid. 9.*

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*"Inner energy*

*levels tend to*

*block the*

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attraction of  
the nucleus fvor  
the valence  
electrons;  
reason why elect  
ronegativity  
goes down in  
value in a  
group/family."-  
is called the  
\_\_\_\_\_.

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: shippo. In

which pair do

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*both compounds  
exhibit*

*predominantly  
ionic bonding?*

*RbCl and CaO.*

*The following  
representation  
of an atom is  
called. A Lewis  
dot structure.*

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e. molar mass of  
 $\text{Fe}(\text{NO}_3)_3 = 241.9$   
 $\text{g}$   $272 \text{ g Fe}(\text{NO}_3)_3$   
 $1 \text{ mol } 241.9 \text{ g} =$   
 $1.12 \text{ mol}$

$\text{Fe}(\text{NO}_3)_3$  f.  
molar mass of  
 $\text{FeSO}_4 = 151.9 \text{ g}$   
 $12.7 \text{ mg FeSO}_4$   $1$   
 $\text{g } 1000 \text{ mg } 1 \text{ mol}$   
 $151.9 \text{ g} = 8.36$   
 $10^{-5} \text{ mol FeSO}_4$

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*g. molar mass of  
LiOH = 23.95 g*

*6.91 103 g LiOH*

*1 mol 23.95 g =*

*288.5 = 289 mol*

*LiOH 14. a.*

*molar mass of*

*CaCO<sub>3</sub> = 100.1 g*

*2.21 10<sup>-4</sup> mol*

*CaCO<sub>3</sub> 100.1 g 1*

*mol = 0 ...*

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order to explore

and make use of

the seemingly

limitless

changes that

matter can

undergo,

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*chemists and  
chemistry*

*students often  
need to answer  
questions that*

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Assessment 3.

The positive

charge of a

group A metal is

equal to its

group number.

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*The charge of a  
group A nonmetal  
is the group  
number minus 8.*

- 4. from the  
number of  
electrons lost*
- 5. -ite or -ate*
- 6.*

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PROBLEMS Write  
the answer on  
the line to the

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Answers  
left. Show all  
your work in the  
space provided.

1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield.
2. 6.0 mol of  $N_2$  are mixed with 12.0

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mol of H  
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