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Mole Homework Packet Answer Key -
Name 1 I= □ Hour ...

Pre AP Chemistry Unit 10 HW Packet
Name _____ 90% of a worksheet must be completed to earn credit for that worksheet! Page 2 of 10 WKS 10.2 □
Calculating with Mass, Moles, & Molecules (1 page) YOU MUST SHOW WORK USING DIMENSIONAL ANALYSIS TO EARN CREDIT FOR THE PROBLEMS! 1. How many grams of ammonia (NH_3)

CHEMISTRY WORKSHEET # 2 MOLE

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PROBLEMS THE MOLE AS A UNIT ...

The mole is a key unit in chemistry. The molar mass of a substance, in grams, is numerically equal to one atom's or molecule's mass in atomic mass units.
Exercises

The Mole - Introductory Chemistry - 1st Canadian Edition

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NAME: HR: _____ PAGE 1

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CHEMISTRY WORKSHEET # 1

MOLAR MASS (GRAM

MOLECULAR/FORMULA WEIGHTS)

We know that grams are actually a measure of the mass of matter and not the weight. Mass is the quantity of matter present; weight is a measure of the pull of gravity on matter and is measured in pounds or newtons.

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For these elements one mole is 6.02×10^{23} molecules. That is, 6.02×10^{23} molecules of hydrogen is one mole of hydrogen, 6.02×10^{23} molecules of nitrogen is one mole of nitrogen, 6.02×10^{23} molecules of oxygen is one mole of oxygen, etc.

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CHEMISTRY WORKSHEET # 5 MOLE PROBLEMS □ MOLAR VOLUME OF A

GAS We have learned that a mole is a mass of material and number of particles. A mole can also be a measure of volume when we are talking about gases.

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Moles(&(Stoichiometry
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CHEMISTRY WORKSHEET # 1
MOLAR MASS (GRAM
MOLECULAR/FORMULA WEIGHTS)

We know that grams are actually a measure of the mass of matter and not the weight. Mass is the quantity of matter present; weight is a measure of the pull of gravity on matter and is measured in pounds or newtons.

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□Molar mass□ is used to describe the mass of one mole of a chemical compound, while □atomic mass□ is used to describe the mass of one mole of an element or the mass of one atom of an element.

Honors Chemistry - Unit 2

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7CHEMISTRY WORKSHEET # 5
MOLE PROBLEMS □ MOLAR VOLUME

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OF A GAS We have learned two definitions of a mol, now we will learn a third. A mole can also be a measure of volume when we are talking about gases.

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The molecular formula = --1 B 13. In an experiment, a student gently heated a hydrated $\text{CuSO}_4 \cdot x \text{H}_2\text{O}$ to remove the water. The following data was recorded:
Mass of empty crucible Mass of crucible contents before heating Mass of crucible contents after heating
Determine the formula of the hydrate. cusc
The formula of the hydrate =

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Stoichiometry Packet - Page 1 of 11
Chemistry Unit 8 ... b. If 6.23 moles of copper are used, how many moles of oxygen must also be used? c. If 83.9

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grams of oxygen react with copper, how many moles of copper (II) oxide will be produced? ... Answer the following g or L ...

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