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calculate the quantity  
listed, in joules, for the  
system of one mole of a gas  
in a cylinder with movable  
piston.

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ones of the following  
reactions are endothermic in  
other words  $\Delta H$  is positive?  
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 $\text{H}_2\text{O}(g)$   $\Delta H$  II.  $2\text{NH}_3 + 22\text{kcal}$

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Example 9.2 The element  
mercury, Hg, is a silvery  
liquid at room temperature.  
The normal freezing point of  
mercury is  $-38.9^{\circ}\text{C}$ , and its  
molar enthalpy of fusion is  
 $\Delta H_{\text{fusion}} = 2.29 \text{ kJ/mol}$ . What

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is the entropy change of the system when 50.0 g of Hg( l) freezes at the normal freezing

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