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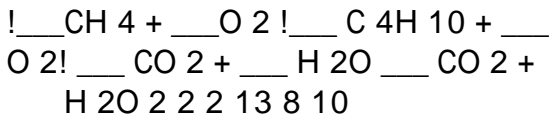
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COMBUSTION REACTIONS! A hydrocarbon and oxygen (O_2) indicate a combustion reaction ! If it is

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combustion, then just write H₂O and CO₂ as products ! Then, balance the equation (can be tricky) ! Examples:



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Products - Solutions Balance the

equations and predict the products for the following reactions: 1) $3 \text{ Na} + 1 \text{ FeBr}_3 \rightarrow 3 \text{ NaBr} + 1 \text{ Fe}$

2) $2 \text{ NaOH} + 1 \text{ H}_2\text{SO}_4 \rightarrow 1 \text{ Na}_2\text{SO}_4 + 2 \text{ H}_2\text{O}$

3) $1 \text{ C}_2\text{H}_4\text{O}_2 + 2 \text{ O}_2 \rightarrow 2 \text{ CO}_2 + 2 \text{ H}_2\text{O}$

4) $1 \text{ NH}_3 + 1 \text{ H}_2\text{O} \rightarrow 1 \text{ NH}_4\text{OH}$

5) 1 PbSO_4

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Products - Solutions Balance the

equations and predict the products for

the following reactions: 1) $3 \text{ Na} + 1$

$\text{FeBr}_3 \rightarrow 3 \text{ NaBr} + 1 \text{ Fe}$ 2) $2 \text{ NaOH} + 1 \text{ H}$

$2 \text{ SO}_4 \rightarrow 1 \text{ Na}_2 \text{ SO}_4 + 2 \text{ H}_2 \text{ O}$ 3) 1 C_2

$\text{H}_4 \text{ O}_2 + 2 \text{ O}_2 \rightarrow 2 \text{ CO}_2 + 2 \text{ H}_2 \text{ O}$ 4) 1

$\text{NH}_3 + 1 \text{ H}_2 \text{ O} \rightarrow 1 \text{ NH}_4 \text{ OH}$ 5) 1 PbSO_4

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Steps to Predicting the Products of Chemical Reactions

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