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3.) The specific heat of a solution is $4.18\text{J}/(\text{g}\cdot\text{C}^\circ)$ and its density is $1.02\text{g}/\text{mL}$. The solution is formed by combining 25.0mL of solution A with 25.0mL of solution B, with each solution initially at 21.4°C . The final temperature of the

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combined solutions is $25.3\text{ }^{\circ}\text{C}$. Calculate the heat of reaction, q_{rxn} , assuming no heat loss to the calorimeter.

6: Types of Chemical Reactions (Experiment) - Chemistry ...

Thermodynamics: Enthalpy of Reaction
and Hess's Law Judy Chen Partner: Mint

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Date: 13 Sept, 2011 Purpose: The purpose of this lab is verify Hess's law by finding the enthalpies of the reactions; NaOH and HCl, NH_2Cl and NaOH, and NH_3 and HCl. The overall enthalpy should equals to the sum of enthalpy of the three reactions in

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Flame Test Lab Questions Answer Key:

Student Laboratory Kit Introduction Just as a fingerprint is unique to each person, the color of light emitted by metals heated in a flame is unique to each metal. In this laboratory activity, the characteristic color of light emitted for calcium, copper, lithium, potassium, sodium, and strontium

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location you will find advance inquiry kits, laboratory kits, an equipment list needed for AP Chem Labs, videos and other resources.

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Lab #1 - Analysis of Food Dyes in Beverages. Lab #2 - Percent Copper in Brass. Lab #3 - Gravimetric Analysis of Calcium in Antacid Tablets. Lab #4 - Acidity of Beverages. Lab #5 - Separation of a Dye Mixture Using Chromatography.

AP Chemistry Lab/Investigations - LHS

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Le Châtelier's Principle

So we did a lab in class and I need help

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finding the final answers. The data table says: the initial mass of $\text{CuCl}_2(\text{g})$ was .52g
Initial mass of aluminum wire (g) was .74
The mass of "leftover" aluminum wire (g) was .68
The actual mass of aluminum reacted (g) was .6
The actual mass of recovered dry copper (g) .22
I need help finding : A) The moles of CuCl_2 B) Moles

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of aluminum that should react ...

AP Chemistry Lab Manual | AP Central — The College Board

Mass of silver nitrate- 1.58g Mass of
copper wire- 2.35g Mass of empty 100 mL
beaker- 50.03g Mass of leftover copper
wire- 2.15g Mass of beaker plus silver

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product- 50.67g 1. Calculate mass and moles of copper wire that reacted in this experiment 2. Calculate the mass and moles of silver metal produced in the reaction 3. Determine the mole ratio-the ratio of the number of moles of silver to ...

I need help on my post lab questions!!? |

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Yahoo Answers

If the vinegar used in the lab had 5.89 g of acetic acid per 100. mL of vinegar, what volume (in mL) of 0.0998 M NaOH would you use to reach the endpoint of the titration? ... Get a free answer to a quick problem. Most questions answered within 4 hours. OR. Find an Online Tutor Now

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Pre-laboratory Assignment: Types of

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Reactions. Many of the reactions use 1 mL of solution. How can you estimate this volume? Do you need to dry the test tubes before using them for the reactions in this experiment? Why or why not? The following reactions are performed, and the results are shown below. Use these results to determine the relative ...

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- Answer the pre-lab questions that appear at the end of this lab exercise. The questions should be answered on a separate (new) page of your lab notebook. Be sure to show all work, round answers,

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and include units on all answers.

Background information can be found in Chapter 15, especially sections 15.6 your textbook (Brown and LeMay).

Pre-lab Questions - Thermodynamics- Enthalpy of Reaction ...

Flinn Advanced Inquiry Laboratory Kits

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for AP * Chemistry Correlated to the College Board Investigative Labs In 2012, the Advanced Placement chemistry curriculum was revised to integrate inquiry, content, and reasoning through an updated set of learning objectives and science practice skills.

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Kinetics of Crystal Violet Fading AP*

Chemistry Big Idea 4, Investigation 11 An
Advanced Inquiry Lab Introduction

Crystal violet is a common, beautiful
purple dye. In strongly basic solutions, the

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bright color of the dye slowly fades and the solution becomes colorless.

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7644 Kinetics of ...

Flame Test Lab Questions Answer Key:

You could readily identify the elements that had obvious colors different from all the others- such as copper that gave off a blue/green color and lithium that gave off a bright red color.

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