

## Online Library Chapter 9 Stoichiometry

# Chapter 9 Stoichiometry

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# Stoichiometry

Stoichiometry - Practice Test ...

Excess Reagent- reactants that are not use up when the reaction is finished  
Limiting Reagent- reagent that is completely used up or reacted  
Percent Yield- amount of product obtained in a chemical reaction\*\*  
To find both of the reagents, you must take both of the reactants and

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find the number of grams in the same compound of the product.

Chapter 9 - Stoichiometry

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questions if you would like ...

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Chapter 9 describes how to use mole ratios, molar masses, conversions, limiting reactants, and percent yield to ...

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From above we can see that if we have  
12.4 mol  $H_2$  we need 4.13 mol  $N_2$ . We

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don't have that much  $N_2$  so the .892 mol of  $N_2$  must be the limiting reagent. We can now determine how much ammonia will be produced using the mole ratio in the balanced equation :

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Modern Chemistry Chapter 9  
Stoichiometry

CHAPTER 9 REVIEW Stoichiometry  
SECTION 3 PROBLEMS Write the

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answer on the line to the left. Show all your work in the space provided.

1. 88%  
The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield.

2. 6.0 mol of  $N_2$  are mixed with 12.0 mol of  $H_2$  according to the following equation:  $N_2(g) + 3H_2(g) \dots$

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## Stoichiometry

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stoichiometry (which you studied in Chapter 3) deals with the mass relationships of elements in compounds. Reaction stoichiometry involves the mass relationships between reactants and products in a chemical

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## Stoichiometry

reaction. Reaction stoichiometry is the subject of this chapter and it is based on

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Chapter 9 - Stoichiometry 9-1

Introduction to Stoichiometry

Composition Stoichiometry - deals with

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## Stoichiometry

mass relationships of elements in  
compounds Reaction Stoichiometry -  
Involves mass relationships between  
reactants and products in a chemical  
reaction I. Reaction Stoichiometry  
Problems A. Four problem Types, One  
Common Solution

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$46/23 \times 1/4 \times 32 = 16$  grams O<sub>2</sub> Modern  
Chemistry Chapter 9 Stoichiometry  
composition stoichiometry deals with the  
mass relationships of elements in  
compounds. reaction stoichiometry  
involves the mass relationships between

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## Stoichiometry

reactants and products in a chemical reaction.

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Chapter 9 focuses on reaction  
stoichiometry: using a balanced chemical  
equation to calculate the number of grams,

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moles, or particles of reactants/products involved in a chemical reaction. Students had an introduction to composition stoichiometry in Chapter 3 and will now move on to some more difficult problems.

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