

Chapter 9 Review Stoichiometry Section 3

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Stoichiometry. SECTION 2. PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. The following equation represents a laboratory preparation for oxygen gas: 2KClO₃(s) → 2KCl(s) + 3O₂(g) ... CHAPTER 9 REVIEW ...

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Chapter 9 describes how to use mole ratios, molar masses, conversions, limiting reactants, and percent yield to ... Stoichiometry Review - ScienceGeek.net Homepage

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From above we can see that if we have 12.4 mol H₂ we need 4.13 mol N₂. We don't have that much N₂ so the .892 mol of N₂ must be the limiting reagent. We can now determine how much ammonia will be produced using the mole ratio in the balanced equation :

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Chapter 9 focuses on reaction stoichiometry, using a balanced chemical equation to calculate the number of grams, moles, or particles of reactants/products involved in a chemical reaction. Students had an introduction to composition stoichiometry in Chapter 3 and will now move on to some more difficult problems.

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