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Chapter 9 focuses on reaction stoichiometry: using a balanced chemical equation to calculate the number of grams, moles, or particles of reactants/products involved in a chemical reaction. Students had an introduction to composition stoichiometry in Chapter 3 and will now move on to some more difficult problems.

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Chapter 9 describes how to use mole ratios, molar masses, conversions, limiting reactants, and percent yield to ...
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From above we can see that if we have 12.4 mol H 2 we need 4.13 mol N 2.We don't have that much N 2 so the .892 mol of N 2 must be the limiting reagent. We can now determine how much ammonia will be produced using the mole ratio in the balanced equation :

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Modern Chemistry 73 Stoichiometry CHAPTER 9 REVIEW Stoichiometry SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. _____ The coefficients in a chemical equation represent the (a) masses in grams of all reactants and products. (b) relative number of moles of reactants and products.

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laboratory preparation for oxygen gas: 2KClO3(s ... CHAPTER 9 REVIEW ...

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mixed with 12.0 mol of H 2 according to the following equation: N 2(g) 3H 2(g) ? 2NH 3(g) N 2; 2.0 mol a.

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