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Structure 8-3

There are two

types of

repulsive forces

between the two

atoms. First, the

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nuclei repel
because they are
both positively
charged.

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Chemical Bonds

Three types: -

Ionic

**Electrostatic
attraction**

between ions

**Covalent Sharing
of electrons**

**Metallic Metal
atoms bonded to
several other
atoms . . .**

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**bond where
electrons from
each atom are
shared**

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**sharing of
electrons occurs
the attachment
between atoms
is called.**

**covalent bond. in
a covalent bond,
the dissociation
energy is
released in the
process of.**

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continued ...

Differentiate

between an

ordinary

covalent bond

and a coordinate

covalent bond.

Give an example

of a molecule

that exhibits

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Programs ... 6A, and 7A) they can accept a complementary number of shared bonds to reach the octet state. Family 4A can share 4 covalent bonds ($4 + 4 = 8$), whereas Families

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**5A, 6A, and 7A
can share 3, 2,
and 1 covalent
bond(s ...**

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When only one

pair of electrons

is shared, such

As in a hydrogen molecule, it is a single covalent bond. The shared electron pair is often referred to as the bonding pair. For a hydrogen molecule, shown in Figure 8.4, each covalently bonded atom equally attracts

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are called:

Intramolecular

Bonds). These

bonds can be a

give-take

relationship

between two

atoms (ionic

bonds) or a

shared electron

relationship

between two

atoms (covalent

bonds).

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