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Chapter 7 Page 1 CHAPTER
SEVEN: QUANTUM THEORY AND
THE ATOM Part One: Light
Waves, Photons, and Bohr
Theory A. The Wave Nature of
Light (Section 7.1) 1.
Structure of atom had been
established as cloud of

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electrons around a heavy
nuclear

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Chapter 7: Electromagnetic
Radiation Radiant Energy ...
Chapter 7 Quantum Theory and
Electronic Structure of
Atoms (Ch7 Chang, Ch7
Jespersen) The Wave Nature
of Light Electromagnetic

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radiation (defined on next slide) can be described as light energy and/or a wave. A wave is a vibrating disturbance by which energy is transmitted.

Chapter 7: Introduction to

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Quantum Theory - Chemistry

...

Chapter 7 Quantum Theory of
the Atom I) Light Waves A)

Nature of Light

Electromagnetic radiation
(or electromagnetic energy
or radiant energy) Types

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visible light x-rays

microwaves radiowaves Travel

as waves electromagnetic

radiation is a result of

oscillating electric and

magnetic fields moving

simultaneously through

space. B) Wave Nature of

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Light

CHAPTER SEVEN QUANTUM THEORY
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Chapter 7: Quantum Theory
and the Electronic Structure
of Atoms. Page 130. Page
129. Title: Chapter 7:

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Chapter 7. Quantum Theory

and Atomic Structure • A

problem arose in

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Rutherford's nuclear model.
A nucleus and electron
attract each other; to
remain apart the electron
must move. • The energy of
the electron's movement must
balance the energy of
attraction. • Physics

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established that a charged
particle moving in a curved

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Quantum Fields: The Real
Building Blocks of the

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Universe - with David Tong -
Duration: 1:00:18. The Royal
Institution Recommended for
you

CHAPTER 7 QUANTUM THEORY AND THE ELECTRONIC STRUCTURE OF ATOMS

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Spell. Test. PLAY. Match.

Gravity. Created by.

vanki004. U of M CHEM 1061.

... Principle Quantum Number

(n) positive integer

(1,2,3...) -indicates

relative SIZE of the orbital

and therefore relative

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distance from the nucleus

Chapter 7 Quantum Theory

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CHAPTER 7: QUANTUM THEORY

AND THE ELECTRONIC STRUCTURE

OF ATOMS 197 7.20 (a) $\lambda = ?$ c

8 7 14 m 3.00 10 s 3.70 10 m

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1 $8.11 \times 10^{-18} \text{ s}^{-1}$? $\lambda = \frac{c}{\nu} = \frac{3.00 \times 10^8 \text{ m/s}}{8.11 \times 10^{14} \text{ s}^{-1}} = 3.70 \times 10^{-7} \text{ m} = 370 \text{ nm}$

(b) Checking Figure 7.4 of the text, you should find that the visible region of the spectrum runs from 400 to 700 nm. 370 nm is in the ultraviolet region of the spectrum. (c) $E = h\nu$.

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Substitute the frequency (?) into this equation to solve for the energy of one quantum

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...

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Tro Chapter 7 Lecture -
Quantum Mechanical Model of
the Atom ... Quantum Theory
Made Easy [1] - Duration:
31:09. ... Chapter 7 -
Periodic ...

Chapter 7. Quantum Theory

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Chapter 7 Quantum Theory

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Chapter 7 Quantum Theory of
the Atom

Chapter 7: Quantum Theory of
the Atom Electromagnetic
Radiation or Radiant Energy:
energy in the form of a wave

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characteristics of a wave:
wavelength, λ - length from
peak to peak or trough to
trough frequency, ν - number
of peaks that pass a set
point per unit time

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