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SPM Form 4

Chemistry Chapter 7 -

Acids and Bases

Chapter 7 – Mixtures of Acids and Bases

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Chapter 7 Acids Bases And

Introduction In Chapter
6, we examined the
equilibrium

concentrations in
solutions of acids and
solutions of bases. In this
chapter, we continue our
discussion of acids and
bases by focusing on the
equilibrium
concentrations of
solutions formed by
mixing acids and bases.

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Bases And

Solutions

Chapter 7: Acids, Bases,

and Solutions Solution –

a homogeneous mixture

Solutions have the same

properties throughout,

containing solute

particles (molecules or

ions) that are too small

to see Solvent – the part

of the solution that is

present in the largest

amount and dissolves

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**Chapter 7 Acids Bases
And**

a substance that can act as either an acid or base (ex. water). An amino acid has a basic amine group and an acidic carboxylic acid group, so it is also amphoteric. Unlike water, amino acids can carry multiple

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charges depending on their environment. At pH 7, most amino acids have a protonated amine group and a deprotonated carboxylic acid group.

Chapter 7 - Acids and Bases - YouTube

Chapter 7 Acid-Base Equilibria 2. Acid-Base Theories • Arrhenius theory (Nobel Prize) An

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acid is any substance that ionizes (partially or completely) in water to give hydrogen ions, H^+ (which associate with the solvent to give hydronium ions, H_3O^+)

$$HA + H_2O \rightleftharpoons H_3O^+ + A^-$$

- A base: any substance that ionizes in water to give hydroxyl ions OH^- .

Chapter 7 ACIDS AND BASES

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Chapter 7 Acids Bases And

Solutions
The p H of a salt
solution obtained by the
reaction between a

strong acid and a strong
base is... [A] is 7 [B] is
less than 7 [C] is more
than 7 [D] cannot be
determined; Sodium

chloride is a salt of...
[A] a weak acid and a
weak base [B] a strong
acid and a strong base
[C] a weak acid and a
strong base [D] a strong

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acid and a weak base

**Chapter 7 - Acids and
Bases Flashcards |
Quizlet**

Chapter 7- Acids and
Bases. STUDY. PLAY.
taste sour, turns litmus
red, reacts with active
metals to form hydrogen
gas, react with bases to
form water and a salt.
Characteristics of acids.
taste bitter, turns litmus

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blue, slippery, reacts
with acids to form water
and a salt.

Answers

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First, according to the
Brønsted-Lowry
definition of acid-base,
acetic acid, $\text{CH}_3\text{CO}_2\text{H}$,
produces H^+ thus it
is an acid. We can write
the equilibrium constant

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Solutions

for the acid is, then K_a

$$K_a = \frac{[\text{H}^+][\text{C}_2\text{H}_3\text{O}_2^-]}{[\text{C}_2\text{H}_3\text{O}_2\text{H}]}$$

$$1.74 \times 10^{-5} = \frac{[\text{H}^+][\text{C}_2\text{H}_3\text{O}_2^-]}{[\text{C}_2\text{H}_3\text{O}_2\text{H}]}$$

$$1.74 \times 10^{-5} = \frac{[\text{H}^+]^2}{[\text{C}_2\text{H}_3\text{O}_2\text{H}]}$$

5 The numerator terms are the product ions

Chapter 7 Acids, Bases, and Solutions

(a) Both acids and bases change colour of all indicators. (b) If an indicator gives a colour change with an acid, it

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Bases And

does not give a change with a base. (c) If an indicator changes colour with a base, it does not change colour with an acid. (d) Change of colour in an acid and a base depends on the type of the indicator.

Chapter 7:

Chapter 7 Mixtures of
Acids and Bases 181 7.2
BUFFERS In Chapter 6,

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Bases And
Solutions
Crossword
Answers

we treated in some detail four types of acid-base solutions: strong acid, strong base, weak acid, and weak base. We now treat buffers, the last type of acid-base solution to be considered. A buffer is defined in the dictionary as a “device that softens the shock of a blow.”

Chapter 7 – Mixtures

Page 15/26

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Bases And
of Acids and Bases

Chapter 7 Acids, Bases,
and Solutions

Instructions: Complete
the crossword puzzle.

Use the clues to help
identify the words. C O
N C E N T R A T E D H
O E I S C A L E Y L U
L O D L T U N S A T U
R A T E D R O R T A R
O I A E L O G D L T S
A T U R A T E D I N D
I C A T O R I N Z V A

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Bases And

SSUPER SATUR
ATED

Crossword

**Chapter 7. Acids,
Bases, and
Equilibrium**

Chapter 7 Acids, Bases,
and Solutions Properties
of Acids and Bases

Litmus is an example of
an indicator, a
compound that changes
color when in contact
with an acid or a base.

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Bases, and Solutions -
nyostrander.us**

35 Chapter 7 ACIDS
AND BASES Exercises

7.2 (a) A nonpolar solvent because phosphorus pentachloride is a nonpolar covalent compound. (b) A polar protic solvent because cesium chloride is ionic.

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SPM - Chemistry -
Form 4 Chapter 7 Acids
and Base - Expected
learning outcomes.

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games, and other study
tools.

Crossword

**NCERT Solutions for
Class 7 Science**

Chapter 5 Acids Bases

...

38 Chapter 7 7.15 (HF

H_2SO_4) + H_2SO_4

(1) ? (H_2F^+ (H_2SO_4)

+ HSO_4^- (H_2SO_4) HF

is acting as a base and H_2F^+

is the conjugate

acid, H_2SO_4 is acting

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as an acid and HSO_4^- is
the conjugate base. 7.17
 HSeO_4^- (base), H_2SO_4
(conjugate acid); H

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gaia_sartorelli includes
24 questions covering
vocabulary, terms and
more. Quizlet

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**Chapter 7 acids and
bases - SlideShare**

Chapter 7:

Neutralization Reactions

. Neutralization

Reaction. A reaction
between an acid and a
base to produce a salt
and water. (Remember
you will need to be able

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Solutions

(to write the ionic and net ionic equations for these reactions.)

Answers

Chapter 7 Acids and Bases - Concept Map

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