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## **Ch 6 - HonorsChemWin s**

Section 6.1 – Ionic  
Bonds When the  
highest occupied  
energy level of an  
atom is filled with  
electrons, the atom



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Answers

is stable and not likely to react. The noble gases have stable electron configurations with eight valence electrons (or two in the case of helium).

**Chapter 6**

**Chemical Bonds**

**Section 6 2**

**Covalent**

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**Bonding ...**

Chapter 6 -

Chemical Bonding

In previous

chapters, students

have studied

subatomic particles

and the properties

of individual atoms.

In Chapter 6, we

will begin studying

how atoms interact

with...

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Section 1-  
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**6 - Section 3**

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Chemical Bonding.

Section 1-

Introduction to

Chemical Bonding.

Objectives: define

chemical bond;

differentiate

between covalent

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and ionic bonding;

Section 4  
explain why

Worksheet  
bonding occurs;

Answers  
use the difference

in electronegativity

to determine

whether a bond is

polar covalent,

nonpolar covalent

or ionic ... Section

3 IONIC BONDING

AND IONIC ...

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Modern Chemistry  
47 Chemical  
Answers  
Bonding CHAPTER  
6 REVIEW Chemical  
Bonding SECTION 4  
SHORT ANSWER

Answer the  
following questions  
in the space  
provided. 1. \_\_\_\_\_  
In metals, the  
valence electrons

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Answers

are considered to be (a) attached to particular positive ions. (c) immobile. (b) shared by all surrounding atoms. (d) involved in covalent bonds.  
2.

## **Chapter 6**

### **Review Chemical Bonding Section 2&3 Flashcards**

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...

Section 1  
Review Previous  
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Bonding CHAPTER  
6 Section 1

Introduction to  
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What is a chemical  
bond and why does  
it form? Section 2  
Covalent Bonding  
and Molecular  
Compounds What  
is a molecular

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Answers

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SECTION REVIEWS

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Ionic Bonding and  
Ionic Compounds  
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terms. mmand.

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3. Ionic Bonding

and Ionic

Compounds.

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Terms in this set

(...) Ionic

compound. a

compound

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Answers

composed of  
positive and  
negative ions that  
are combined so  
that the numbers  
of positive ...

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**REVIEW Chemical**  
**Bonding**

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Answers

## **Section 6.1 6.1**

### **Ionic Bonding**

CHAPTER 6 REVIEW

Chemical Bonding

SECTION 6-2

SHORT ANSWER

Answer the

following questions

in the space

provided. 1. Use



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Answers

the concept of potential energy to describe how a covalent bond forms between two atoms. 2. Name two elements that form compounds that are exceptions to the octet rule.

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Section 4  
Class Date Section

Quiz: Metallic

Bonding In the

space provided,

write the letter of

the term or phrase

that best

completes each

statement or best

answers each

question. 1.

Chemical bonding

in metals is a. the

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Answers

same as ionic bonding. b. the same as covalent bonding. c. a combination of ionic and covalent bonding.

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chapter 6  
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bonding  
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Chapter 6:

Section 4  
Chemical Bonding,  
Modern Chemistry.

Answers  
This is a working presentation of the notes for this chapter. Meaning that we may or may not cover all of the material here.

**Chapter 6**

**Sections 1-5**

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Section 4  
**Review sheets -  
Key - Course**

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Chapter 6 Section

4 Section 6-4:

Metallic Bonding

6-4-3 Explain why

metals are

malleable and

ductile but ionic-

crystalline

compounds are

not. The metallic

bond is the same in

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Worksheet  
Answers

all directions  
throughout the  
metallic structure  
allowing the atoms  
to slide past each  
other. This sliding  
is why metals are  
ductile and  
malleable. Ionic  
compound

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**Chemical**  
**Bonding**

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## **Chapter 6**

### **Chemical**

### **Bonding Section**

CHAPTER 6 REVIEW

Chemical Bonding

SECTION 1 SHORT

ANSWER Answer

the following

questions in the

space provided. 1.

a A chemical bond

between atoms

results from the

attraction between



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Answers

the valence electrons and of different atoms. (a) nuclei (c) isotopes (b) inner electrons (d) Lewis structures

2. b A covalent bond consists of (a) a shared electron.

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