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Chapter Four

[Arrangement of

Electrons in Atoms]

CHEMISTRY

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Chapter 4

Arrangement Of CHAPTER 4.

(Arrangement of
Electrons) The lowest
energy state of an atom
is its ground state.

(usually it's the lowest
levels) A photon is a
particle of
electromagnetic
radiation having zero
mass and carrying a
quantum of energy.

When a photon strikes a
atom it gives the atoms

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more energy.
Electrons In

Chapter 4
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electrons Chemistry
Bishop ...

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Arrangement of
Electrons in Atoms
CHAPTER 4 REVIEW

Arrangement of
Electrons in Atoms
Teacher Notes and
Answers Chapter 4

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Chapter 4

Arrangement Of

SECTION 1 SHORT

ANSWER 1. In order

for an electron to be

ejected from a metal

surface, the electron

must be struck by a

single photon with at

least the minimum

energy needed to knock

the electron loose. 2.

CHEMISTRY

CHAPTER 4

(Arrangement of

Page 8/30

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Arrangement Of
Electrons)

Accelerated Chemistry

1 Arrangement of

Electrons in Atoms

CHAPTER 4

Arrangement of

Electrons in Atoms

SECTION 1 The

Development of a New

Atomic Model

OBJECTIVES 1.

Explain the

mathematical

relationship among the

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Arrangement Of
speed, wavelength, and
frequency of
Electrons In
electromagnetic
radiation. 2.

CHAPTER 4

Arrangement of Electrons in Atoms

Modern Chemistry 5

Arrangement of

Electrons in Atoms

CHAPTER 4 REVIEW

Arrangement of

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Chapter 4

Arrangement Of

SECTION 2 SHORT
ANSWER Answer the

following questions in
the space provided. 1.

_____ How many
quantum numbers are
used to describe the
properties of electrons
in atomic orbitals? (a) 1
(c) 3 (b) 2 (d) 4 2.

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Arrangement of
Electrons Flashcards**

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Arrangement Of

...

Electrons In

Atoms Test

Answers

4. c How many

electrons can an energy

level of $n = 2$ hold? (a)

32 (c) 8 (b) 24 (d) 6 . 5.

c Compared with an

electron for which $n = 2$,

an electron for which n

$= 4$ has more (a) spin.

(c) energy. (b) particle

nature. (d) wave nature.

6. c . According to

Bohr, which is the point

in the figure below

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where electrons cannot
reside? (a) point A (c)
point C

Answers

Chapter 4

Arrangement Of Electrons

Hund's rule. orbitals of
equal energy are each
occupied by one
electron before any
orbital is occupied by a
second electron, and all

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Arrangement Of

electrons in singly

occupied orbitals must

have the same spin.

Pauli's exclusion

principle. no two

electrons in an atom can

have the same set of

quantum numbers.

4 Arrangement of Electrons in Atoms

Chapter 4 -

Arrangement of

Electrons in Atoms. The

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Chapter 4

Arrangement Of
“Puzzle” of the nucleus:

Electrons In
Atoms Test
Answers
• Protons and electrons
are attracted to each

other because of
opposite charges •

Electrically charged
particles moving in a
curved path give off
energy • Despite these

facts, atoms don't
collapse. 4-1 The
Development of a New
Atomic Model.

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yazvac

Chapter Four

[Arrangement of

Electrons in Atoms]

Chapter Five [The

Periodic Law] Chapter

Six [Chemical Bonding]

Chapter Seven

[Chemical Formulas and

Chemical Compounds]

... Chapter Lab - Flame

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Chapter 4

Arrangement Of

tests . Homework:

Section 1: Chapter
review 1 thru 14.

Section 2: Chapter
review 15 thru 22.

Chapter 4 Test

**Review: Arrangement
of Electron in Atoms ...**

About This Chapter.

The Arrangement of

Electrons in Atoms

chapter of this Holt

McDougal Modern

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Arrangement Of

Chemistry Companion
Electrons In
Course helps students

learn the essential

lessons associated with

how electrons are

arranged in atoms. Each

of these simple and fun

video lessons is about

five minutes long and is

sequenced to align with

the Arrangement...

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REVIEW

Arrangement

of Electrons in Atoms

100 CHAPTER 4

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Answers

FIGURE 4 Excited neon atoms emit light when electrons in higher energy levels fall back to the ground state or to a lower-energy excited state.

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Arrangement of

Electrons in Atoms

ARRANGEMENT OF

ELECTRONS IN

ATOMS 91 SECTION

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Chapter 4

Arrangement Of 4-1 OBJECTIVES

Explain the
mathematical

relationship among the speed, wavelength, and frequency of electromagnetic radiation. Discuss the dual wave-particle nature of light. Discuss the significance of the photoelectric effect and the line-emission spectrum of hydrogen to

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Arrangement Of
the development of the
Electrons In
atomic model.

Atoms Test

CHAPTER 4

REVIEW

Arrangement of

Electrons in Atoms

CHAPTER 4 REVIEW

Arrangement of

Electrons in Atoms

SECTION 3 SHORT

ANSWER Answer the
following questions in
the space provided. 1.

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Answers

State the Pauli exclusion principle, and use it to explain why electrons in the same orbital must have opposite spin states. The Pauli exclusion principle states that no two electrons in an atom may have the

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Chapter 4

Chapter 4 Practice Test
(Enriched):

Electromagnetic

Spectrum: Wavelength
and frequency:

Photoelectric effect:

Hydrogen's Line

emission spectrum:

Photon emission and
absorption: Electrons
accommodated in energy

levels and sublevels:

Quantum numbers of
the first 30 atomic

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Arrangement Of

orbitals: Relative

energies of orbitals A:

Writing electron

configurations: Relative

energies of orbitals B:

Orbital notation for

three noble gases:

Orbital notation for

Argon and Potassium:

Supplements from

Chapter:

Holt McDougal

Modern Chemistry

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Chapter 4:

Arrangement of ...

Chapter 4-Test

Arrangement of

Electrons in Atoms In

the previous chapter,

basic atomic structure

was introduced and

nuclear chemistry was

reviewed. In this

chapter, we will study

how electrons are...

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