

Chapter 22 Hydrocarbons

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SECTION 22.1 HYDROCARBONS(pages 693-701)

718 Chapter 22 22CHAPTER Study Guide Key Concepts 22.1 Hydrocarbons • Because carbon has four valence electrons, car-bon atoms always form four covalent bonds. • The carbon atoms in an alkane can be arranged in a straight chain or in a chain that has branches. • Molecules of hydrocarbons, such as alkanes, are nonpolar molecules. 22.2 ...

Chapter 22: Substituted Hydrocarbons and Their Reactions

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AP Chemistry Chapter 22 - Organic and Biological Molecules 22.1 Alkanes: Saturated Hydrocarbons A. Straight-chain Hydrocarbons 1. Straight-chain alkanes have the formula C_nH_{2n+2} 2. Carbons are sp^3 hybridized The First 10 Alkanes # of Carbons Name Formula (C_nH_{2n+2}) 1 Methane CH₄ 2 Ethane C₂H₆ 3 Propane C₃H₈ 4 Butane C₄H₁₀ 5 Pentane C₅H₁₂ ...

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Chapter 22 Hydrocarbon Compounds; Shared Flashcard Set. Details. Title. Chapter 22 Hydrocarbon Compounds. Description. Key Concepts and Vocabulary. Total ... an atom or group of atoms that can take the place of a hydrogen atom on a parent hydrocarbon molecule. Term. Alkyl group: Definition. a hydrogen substituent. it can be one carbon or ...

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Chapter 22 - Chapters 15-22 (No 18)

Chapter 14 - Gases; Chapter 15 - Solutions; Chapter 16 - Energy and Chemical Change; Chapter 17 - Reaction Rates; Chapter 18 - Equilibrium; Chapter 19 - Acids and Bases; Chapter 20 - Redox Reactions; Chapter 21 - Electrochemistry; Chapter 22 - Hydrocarbons; Chapter 23 - Substituted Hydrocarbons and Their Reactions; Chapter 24 - The Chemistry of ...

Chapter 22 Hydrocarbons

You are at chapter 22!! Let's goooo!!!! First of all, a hydrocarbon is the simplest organic compound and it only contains carbon and hydrogen. For this, we can use our background knowledge to know that since carbon has four valence electrons, it can form four covalent bonds.

Chapter 22 Study Guide For Content Mastery Hydrocarbons ...

Chapter 22 Cracking - process by which hydrocarbons are broken down into smaller molecules Petroleum - fossil fuel containing straight and branched chain alkanes Distillation - process by which petroleum is separated into fractions Lignite - Brown coal, consisting of about 50% carbon

The First 10 Alkanes # of Carbons Name Formula H_{2n+2}

Chapter 22 Hydrocarbon Compounds 241 position of their atoms a double bond at least one substituent on each carbon of the double bond cis-2-butene symmetrical superimposed true It is a carbon with four different atoms or groups attached. They contain an asymmetric carbon atom.

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Title: Chapter 22 Hydrocarbon Compounds 1 Chapter 22 Hydrocarbon Compounds kerosene Diesel fuel gasoline 2 Organic Chemistry Hydrocarbons Scientists used to think that only living organisms could synthesize the carbon compounds found in their cells. They called the study of these compounds organic chemistry. However,

Chapter 22: Hydrocarbon Compounds - Chemistry by Anna

786 Chapter 22 • Substituted Hydrocarbons and Their Reactions Section 22.12.1 Figure 22.1 All of these items contain at least one of the functional groups that you will study in this chapter. For example, the fruit and flowers have sweet-smelling aromas that are due to ester molecules. Objectives Define functional group, and give examples.

Chapter 22 Hydrocarbon Compounds Flashcards

Chapter 22 Hydrocarbon Compounds 241 position of their atoms a double bond at least one substituent on each carbon of the double bond cis-2-butene symmetrical superimposed true It is a carbon with four different atoms or groups attached. They contain an asymmetric carbon atom.

CHAPTER 22 Study Guide - Quia

This unit looks at aliphatic and aromatic hydrocarbons, including alkyl substituted hydrocarbons and structural and geometric isomers. There is also a brief introduction to functional groups. Chapter 22 - Hydrocarbon Compounds

Chapter 22 Hydrocarbons Study Guide Answers

Chapter 22 Hydrocarbon Compounds. Flashcard maker : Paula Corcoran. How is the number of valence electrons in carbon atoms related to the bonds that carbon atoms form? because carbon has four valence electrons, a carbon atom always forms four covalent bonds.

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atom. Therefore, the simplest hydrocarbon molecule consists of a carbon atom bonded to four hydrogen atoms, CH₄. This substance is called methane and is an excellent fuel and the main component of natural gas. 698 Chapter 22 Hydrocarbons Figure 22-1 Carbon is a nonmetal element located in group 4A of the periodic table. Elements in this

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