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ANSWERS. 1.

Determine what is oxidized and what is reduced in each reaction. Identify the oxidizing agent and the reducing agent, also.

1. $2\text{Sr} + \text{O}_2 \rightarrow 2\text{SrO}$
Sr 0 to Sr^{2+} ; oxidized/reducing agent
 O^0 to O^{2-} ; reduced/ox. ag.

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edition ... redox
reaction—from higher
to lower potential
energy.

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Potential Difference •

The energy that would
be required to move a
unit of electrical
charge from ...

Answer: -0.40 V .

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redox couples. such as $\text{Zn} / \text{Zn}^{2+}$ and $\text{Fe}^{2+} / \text{Fe}^{3+}$ (reduced form / oxidized form) when characterizing redox reactions. (Recall the . half-reaction . method of balancing redox equations). In this exercise we deal with . electrochemical

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cells, in which two redox couples are placed in separate compartments as opposed to their direct contact. Cell ...

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Redox Reactions

Figure 20-1 The

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reaction of magnesium and oxygen involves a transfer of electrons from magnesium to oxygen. Therefore, this reaction is an oxidation–reduction reaction. Using the classifications given in Chapter 10, this redox reaction also is classified as a combustion reaction.

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X X X X X X X X 2Mg
2MgO Each ...

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the reducing agent,
also. 1. $2\text{Sr} + \text{O}_2$ 2.

$2\text{Li} + \text{S}$. 2SrO Li_2S . 3.
 $2\text{Cs} + \text{Br}_2$

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ANSWERS I.

Determine what is
oxidized and what is
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reaction. Identify the oxidizing agent and the reducing agent, also. 2SrO Sr^0 to Sr^{2+} ;

oxidized/reducing agent O^0 to O^{2-} ;

reduced/ox. ag. 1. $2\text{Sr} + \text{O}_2$ 2. $2\text{Li} + \text{S}$ Li_2S
 Li^0 to Li^{1+} ;

oxidized/red. ag.

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