

Chapter 18 Review Chemical Equilibrium

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In this video I'll explain dynamic chemical equilibrium and teach you how to generate an equilibrium constant expression, K_c , for an equilibrium chemical reaction. I'll also teach you how to ...

Chemistry Chapter 18 Review Chemical Equilibrium ...

Modern Chemistry 149 Chemical Equilibrium CHAPTER 18 REVIEW Chemical Equilibrium SECTION 4 SHORT ANSWER Answer the following questions in the space provided. 1. Match the solution type on the right to the corresponding relationship between the ion product and the K_{sp} for that solution, listed on the left. ____ The ion product exceeds the K_{sp} .

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sometimes referred to as the chemical-equilibrium expression. The H_2 , I_2 , HI Equilibrium System Consider the reaction between H_2 and I_2 vapor in a sealed flask at an elevated temperature. The rate of reaction can be followed by observing the rate at which the violet color of the iodine vapor diminishes, as $[C]^x[D]^y [A]^n[B]^m$ 556 CHAPTER 18

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CHAPTER 18 REVIEW Chemical Equilibrium SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. Write the equilibrium expression for the following hypothetical equation: $3A(aq) + B(aq) \rightleftharpoons 2C(aq) + 3D(aq)$ 2. a. Write the appropriate chemical equilibrium expression for each of the following equations. Include the value of K ...

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CHAPTER 18 Chemical Equilibrium

, 1. In a bottle of unopened cola, the CO_2 gas dissolved in the liquid is in equilibrium with the CO_2 gas above the liquid. The dissolved gas reacts with water molecules in the cola to form carbonic acid, which also dissociates into carbon dioxide and water. Which chemical equation(s) best describe this equilibrium system?</br> a.

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