

Chapter 17 Thermochemistry Study Guide Answers

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Study Guide Thermochemistry $Q = m C \Delta T$ $\Delta H = -q$ $C = 4.184 \text{ J moles g } ^\circ\text{C}^{-1}$. When 150-g sample of KCl dissolves in 65.0 g of water in a calorimeter, the temperature

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CHAPTER 17, Thermochemistry (continued) 2. What is calorimetry? Calorimetry is the accurate and precise measurement of heat change for chemical and physical processes. 3. Use Figure 17.5 on page 511. Circle the letter next to each sentence that is true about calorimeters. G The calorimeter container is insulated to minimize loss of heat to or

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Chemistry Chapter 17: Thermochemistry: Study Guide: STUDY. Flashcards. Learn. Write. Spell. Test. PLAY. Match. Gravity. Created by. Caleb_Tucker1. Terms in this set (55) Energy. The capacity for doing work or supplying heat. Thermochemistry. The study of energy changes that occur during chemical reactions and changes in state.

SECTION 17.1 THE FLOW OF ENERGY HEAT AND WORK (pages 505–510)
CHAPTER 17, Thermochemistry (continued) GUIDED PRACTICE PROBLEM 12 (page 513) 12. When 50.0 mL of water containing 0.50 mol HCl at 22.5°C is mixed with 50.0 mL of water containing 0.50 mol NaOH at 22.5°C in a calorimeter, the temperature of the solution increased to 26.0°C.

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Chapter 17 Thermochemistry 183 ... 184 Guided Reading and Study Workbook CHAPTER 17, Thermochemistry (continued) 6. In thermochemical calculations, is the direction of heat flow given from the ... thermochemical equation in the first paragraph on page 517 as a guide. SECTION 17.3 HEAT IN CHANGES OF STATE (pages 520–526)

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Name ____ Date ____ Period ____ Chapter 17 Thermochemistry Study Guide 17.1 – 17.2 Thermochemical Equations 1. Make the following conversions: a. 444 cal to joules = 1.86 x 10³ J b. 1.8 kJ to joules = 1.8 x 10³ J c. 0.45 kJ to calories = 1.1 x 10² cal 2. Classify each of these processes as endothermic or exothermic: a. condensing steam ...

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