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Question 6 The following graph shows the stability of a molecule as its structure is varied during conformational analysis. If energy minimisation was carried out on a starting structure appearing as shown on the graph, at which point would energy minimisation stop?

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1 . The instantaneous rate is the rate of a reaction at any particular point in time, a period of time that is so short that the concentrations of reacta

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Major topics: solubility product expressions (Ksp), solubility/Ksp calculations, common ion effect calculations, pH's effect on solubility, & quantitative precipitation (Q vs Ksp)

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Chapter 17: Alcohols and Phenols phenol (aromatic alcohol) pKa- 10 alcohol pKa- 16-18 O C H C O CC H enol keto chemistry dominated by the keto form CO H sp3 O H Alcohols contain an OH group connected to a saturated carbon (sp3) Phenols contain an OH group connected to a carbon of a benzene ring 77 O H H RO R'

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Chapter 17 – Additional Aspects of Aqueous Equilibria: Part 4 of 21

Major topics: common ion effect, definition of a buffer, pH of a buffer calculations (Henderson-Hasselbach), & predicting reactants of a buffer + acid/base.

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Chapter 17 a ntroduCtion o B synthetiC p

In basic solution, [OH ?] > 1 × 10 ?? M > [H +].Hydrogen ion cannot appear as a reactant because its concentration is essentially zero. If it were produced, it would instantly react with the excess hydroxide ion to produce water.

Chapter 17 The Chemistry Of

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In this lecture I'll teach you how to determine and calculate how the common ion effect affects solubility of a saturated solution. I'll also show how pH cha...

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