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Chapter 15 Water and Aqueous Systems Worksheet Answers ...
1. Chapter 15 " Water and Aqueous Systems " Pre-AP Chemistry Charles Page High School Stephen L. Cotton 2. Section 15.1 Water and it ' s Properties OBJECTIVES: – Explain the high surface tension and low vapor pressure of water in terms of the structure of the water molecule and hydrogen bonding. 3.

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The high surface tension of water is due to the: a. small size of water molecules. b. low mass of water molecules. c. hydrogen bonding between water molecules. d. covalent bonds in water molecules. ____ 12. Salts and other compounds that remove moisture from air are said to be: a. efflorescent. c. colloidal. b. surfactant. d. hygroscopic. 10 9 ...

Chapter 15 - Water and Aqueous Systems - 15.3 Heterogeneous ...

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Chapter 15 - Water and Aqueous Systems - 15.1 Water and Its Properties - 15.1 Lesson Check - Page 493: 1 Answer The hydrogen bonding in water cause the high surface tension and low vapor pressure of water.

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WATER AND AQUEOUS SYSTEMS chapter ... Crystals of copper sulfate pentahydrate always contain five molecules of water for each copper and sulfate ion pair.

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Chapter 15 - Water and Aqueous Systems - 15 Assessment - Page 511: 64 Answer Anions are attracted to the hydrogen atoms of the water molecule, because hydrogen atoms are slightly positive.

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Chapter 15 Water and Aqueous Systems 159 SECTION 15.1 WATER AND ITS PROPERTIES (pages 445 – 449) This section describes the properties of water in the liquid and solid states and explains how hydrogen bonding affects the surface tension and vapor pressure

05 CTR ch15 7/12/04 8:14 AM Page 387 WATER AND AQUEOUS ...

Chapter 13 - States of Matter Chapter 14 - Behavior of Gases Chapter 15 - Water and Aqueous Systems Chapter 16 - Solutions Chapter 17 - Thermochemistry Chapter 18 - Reaction Rates and Equilibrium Chapter 19 - Acids, Bases and Salts Chapter 20 - Oxidation-Reduction Reactions

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Quia - Chapter 15 "Water and Aqueous Systems"

Chapter 15 - Water and Aqueous Systems - 15.3 Heterogeneous Aqueous Systems - 15.3 Lesson Check - Page 507: 21 Answer Colloids cannot be separated through filtering because their particles are smaller than those of suspensions and they do not settle out over time.

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Chapter 15 (Water and Aqueous Systems) Test Study Guide The bonds between the hydrogen and oxygen atoms in a water molecule are polar covalent bonds. The covalent bonds are polar because the oxygen atom has a greater electronegativity than the hydrogen atoms. The bonds between adjacent water molecules are called hydrogen bonds.

SECTION 15.1 WATER AND ITS PROPERTIES (pages 445 – 449)

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