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Chapter 12

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In Example 12.2.1 and Example 12.2.2, the identity of the limiting reactant has been apparent: $[\text{Au}(\text{CN})_2]^-$, LaCl_3 , ethanol, and para-nitrophenol. When the limiting reactant is not apparent, we can determine which reactant is limiting by comparing the molar amounts of the reactants with their coefficients in the balanced chemical equation, just as we did in Section 11.4 .

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CHAPTER 11 SOLUTIONS MANUAL
Section 11.1 Defining
Stoichiometry pages 368–372
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371–372 1. Interpret the
following balanced chemical
equations in terms of
particles, moles, and mass.
Show that the law of
conservation of mass is

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Chapter 12 Multiple Choice
Identify the letter of the
choice that best completes
the statement or answers the
question. ____ 1. The first
step in most stoichiometry
problems is to ____.

- a. add the coefficients of the reagents
- c. convert given quantities to volumes
- b. convert given quantities to moles
- d. convert given quantities to masses

____ 2.

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of Equations Using Balanced
Chemical Equations Chemists
use balanced chemical
equations as a basis to

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calculate how much reactant is needed or product is formed in a reaction.

Stoichiometry = calculation of quantities in chemical reactions is a subject of chemistry.

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