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Methods "This"ll only take a few seconds," said the medical technician as he swabbed Emilio's middle finger. The technician took a small lancet and pricked the fingertip. Emilio watched as a small drop of blood formed.

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special dipole-dipole attractions that occur between molecules that contain a hydrogen atom and bonded to a small, highly electronegative atom with at least one lone pair of electrons. Viscosity. measure of the resistance to flow.

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12 States of Matter -
Mr. McKnight

Clawson High School

Section 12 .2 Forces

of Attraction. 1.

intermolecular forces.

2. intermolecular

forces. 3. Dispersion

forces are weak

forces that result from

temporary shifts in the

density of electrons in

electron clouds.

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Chemistry Chapter 12

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In a balanced equation, the ratio between the numbers of moles of a reactant that remains after a chemical reaction stops and a reactant that is totally consumed during a chemical reaction describes the behavior of matter in terms of particles in

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moti Intramolecular
or intermolecular
forces.

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13. with formulas.

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all matter is made up of small particles in constant, random motion, having billions of elastic collisions per second

Temperature average kinetic energy of a substance; as this goes up, Energy (& speed) of particles increases

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□ Identify the limiting reactant in a chemical equation. □ Identify the excess reactant and calculate the amount remaining after the reaction is complete.

States of Matter

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Change. viscosity. a

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Change
measure of the
resistance of a liquid
to flow, which is
affected by the size
and shape of
particles, and
generally increases
as the temperature
decreases and as
intermolecular forces
increase.

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chemistry. □ Describe oxidation and reduction (redox) reactions. □ Discuss the different forms of matter: elements, compounds, and mixtures. □ Explain the difference between solutions, suspensions, and emulsions. □ Explain pH and the pH scale.

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Atmosphere equal to
760 mmHg or
101.3kPa.

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Viscosity.

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Chemistry: The Molecular Nature Of

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MANUAL 30.

Compare and contrast sublimation and evaporation. In both processes, the substances become a vapor. During sublimation, the substance goes from the solid phase directly to the vapor phase. During

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evaporation, particles

in a liquid gain

enough

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