

Chapter 11 Chemical Reactions Answers

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oxidation-reduction reactions A chemical reaction that exhibits a change in the oxidation states of one or more elements in the reactants that has the general form oxidant + reductant \rightarrow reduced oxidant + oxidized reductant. The general forms of these five kinds of reactions are summarized in Table 11.6.1, along with examples of each.

Chapter 11: Chemical Reactions

Write correct formulas for the products in these double replacement reactions. $\text{H}_3\text{PO}_4 + 2 \text{K}_2\text{CO}_3 + \text{BaCl}_2 \rightarrow \text{Ba}_3(\text{PO}_4)_2 + 6 \text{KCl}$
5) $\text{AgNO}_3 + \text{KCl} \rightarrow \text{AgCl} + \text{KNO}_3$
6) $\text{H}_2\text{SO}_4 + \text{Ca} \rightarrow \text{CaSO}_4 + \text{H}_2$
7) $\text{BCl}_3 + \text{H}_2\text{O} \rightarrow \text{H}_3\text{BO}_3 + 3 \text{HCl}$
8) $\text{P}_4\text{O}_{10} + 6 \text{H}_2\text{O} \rightarrow 4 \text{H}_3\text{PO}_4$
9) $\text{Ag}_2\text{C}_2\text{H}_3\text{O}_2 + \text{K}_2\text{CrO}_4$

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Chapter 11: Chemical Reactions Study Guide $\square\square$ chemical equation A representation of a chemical reaction with reactants on the left, products on the right, and an arrow separating the two skeleton

CHAPTER 11 CHEMICAL REACTIONS GUIDED READING ANSWER KEY ...

Section 11.3 – Reactions in Aqueous Solution. A complete ionic equation is an equation that shows dissolved ionic compounds as dissociated free ions. A spectator ion is an ion that is not directly involved in a reaction. Compounds that are aqueous will break into ions, and compounds that are solid will remain bonded.

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