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Chemistry

Writing Formulas (KEY) (Criss-Cross Method) Cl-CO 3-2  
OH-SO 4-2 PO 4-3 NO 3-Na<sup>+</sup> NaCl Na 2CO 3 NaOH Na  
2SO 4 Na 3PO 4 NaNO 3 NH 4 + NH 4Cl (NH 4) 2CO 3 NH  
4OH (NH 4) 2SO 4 (NH 4) 3PO 4 NH 4NO 3 K<sup>+</sup> KCl K 2CO 3  
KOH K 2SO 4 K 3PO 4 KNO 3 Ca<sup>+2</sup> CaCl 2 CaCO 3  
Ca(OH) 2 CaSO 4 Ca 3(PO 4) 2 Ca(NO 3) 2 Mg<sup>+2</sup> MgCl 2  
MgCO 3 Mg(OH) 2 MgSO 4 Mg 3(PO 4) 2 Mg(NO 3) 2 Zn<sup>+2</sup>  
ZnCl 2 ZnCO 3 Zn(OH) 2 ZnSO 4 Zn 3(PO 4) ...

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r i c Page 4:2 ANSWER KEY WRITING FORMULAS  
(CRISS-CROSS METHOD) Name IONIC BONDING Name \_  
IonicbondingoccurswhenametcI  
hcnstersonemoreelectrons10anonmelolln on ...

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I The answer is expressed as 255.5 mL since 46.0 mL has only one decimal place. 11 Perform the following operations expressing the answer in the correct number

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of significant figures. 1, 1.35 m x 2.467 m = Chem\* IF8766  
10 Olnstructional Fair, Inc. 60

SIGNIFICANT FIGURES Name

Word Equations Answer Key 1. Zinc and Lead (II) nitrate  
react to form Zinc Nitrate and Lead.  $Zn + Pb(NO_3)_2$

$Zn(NO_3)_2 + Pb$  Single Replacement 2. Aluminum

Bromide and Chlorine gas react to form Aluminum

Chloride and Bromine gas.  $2AlBr_3 + 3Cl_2 \rightarrow 2AlCl_3 + 3Br_2$

2 Single Replacement 3.

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figures name.

Balancing Chemical Equations Answer Key

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FIGURES Name ... The answer is expressed as 190,000 cm<sup>3</sup> since 19 cm has only two significant figures. When adding and subtracting, limit and round your answer to the least number of decimal places in any of the numbers that make up your answer.

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Chemistry IF8766 62 . STOICHIOMETRY: MASS-MASS PROBLEMS 1.2KC103 + \_302 Name How many grams of potassium chloride are produced if 25 g of potassium chlorate decompose? / 25 ken How many grams of hydrogen are necessary to react completely with 50.0 g of nitrogen in the above reaction?

KMBT 654-20140114144911 - Berger's Chemistry Class  
Balancing Chemical Equations -Answer Key Balance the equations below: 1)  $1 \text{ N}_2 + 3 \text{ H}_2 \rightarrow 2 \text{ NH}_3$  2)  $2 \text{ KClO}_3 \rightarrow 2 \text{ KCl} + 3 \text{ O}_2$  3)  $2 \text{ NaCl} + 1 \text{ F}_2 \rightarrow 2 \text{ NaF} + 1 \text{ Cl}_2$  4)  $2 \text{ H}_2 + 1 \text{ O}_2 \rightarrow 2 \text{ H}_2\text{O}$  5)  $1 \text{ Pb(OH)}_2 + 2 \text{ HCl} \rightarrow 2 \text{ H}_2\text{O} + 1 \text{ PbCl}_2$  6)  $2 \text{ AlBr}_3 + 3 \text{ K}_2\text{SO}_4 \rightarrow 6 \text{ KBr} + 1 \text{ Al}_2(\text{SO}_4)_3$

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Chemistry IF8766 . Name THE TRIPLE AND FOUR BEAM

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10 100 50 0.5 ZOO too 60 600 30 ... Chemistry IF8766 .  
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